



**55. Phosphorus has three allotropic forms — (i) white phosphorus (ii) red phosphorus and (iii) black phosphorus. Write the difference between white and red phosphorus on the basis of their structure and reactivity.**

**Solution:**

1. White phosphorus

(i) It is soft waxy solid having garlic odour

(ii) It is poisonous

(iii) Has low melting point and boiling point because P<sub>4</sub> molecules are held together by weak van der Waals forces of attraction

2. Red Phosphorus

(i) It is hard, crystalline, odourless solid

(ii) It is non-poisonous

(iii) It has a high melting point because of its polymeric structure.

**56. Give an example to show the effect of concentration of nitric acid on the formation of oxidation product.**

**Solution:**



**57. PCI<sub>5</sub> reacts with finely divided silver on heating and a white silver salt is obtained, which dissolves on adding excess aqueous NH<sub>3</sub> solution. Write the reactions involved to explain what happens.**

**Solution:**

PCI<sub>5</sub> reacts with silver to form white silver salt (AgCl). Which then dissolves on adding excess aqueous NH<sub>3</sub> to form a soluble complex



**58. Phosphorus forms a number of oxoacids. Out of these oxoacids phosphinic acid has strong**

# Glencoe Chemistry Chapter 5 Assessment Answers

**Anthony S. Fauci**



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