# Nuclear Chemistry

#### **SECTION 1**

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- 1. \_\_\_\_\_ Based on the information about the three elementary particles in the text, which has the greatest mass?
  - (a) the proton
  - (b) the neutron
  - (c) the electron
  - (d) They all have the same mass.
- 2. \_\_\_\_ The force that keeps nucleons together is
  - (a) a strong nuclear force.
  - (b) a weak nuclear force.
  - (c) an electromagnetic force.
  - (d) a gravitational force.
- The stability of a nucleus is most affected by the
  - (a) number of neutrons.
  - (b) number of protons.
  - (c) number of electrons.
  - (d) ratio of neutrons to protons.
- If an atom should form from its constituent particles,
  - (a) matter is lost and energy is taken in.
  - (b) matter is lost and energy is released.
  - (c) matter is gained and energy is taken in.
  - (d) matter is gained and energy is released.
- For atoms of a given mass number, those with greater mass defects, have
  - (a) smaller binding energies per nucleon.
  - (b) greater binding energies per nucleon.
  - (c) the same binding energies per nucleon as those with smaller mass defects.
  - (d) variable binding energies per nucleon.
- 6. Based on Figure 1.1 of the text, which isotope of He, helium-3 or helium-4,
  - a. has the smaller binding energy per nucleon?
  - b. is more stable to nuclear changes?
- The number of neutrons in an atom of magnesium-25 is \_\_\_\_\_\_.
- Nuclides of the same element have the same \_\_\_\_\_\_.

# **Glencoe Chemistry Science Not Answer Key**

Geoffrey A Ozin, André Arsenault

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