mudtiply reduction rx. by 3: 3[ClO, (aq) + 8H*(aq) + 8e --- $CI^{-}(aq) + 4H_2O(\ell)$ 3ClO₄"(aq) + 24H*(aq) + 24e" --- $3CI^{-}(aq) + 12H_{2}O(\ell)$ colored block $4Br^{-}(aq) + 12H_{2}\Theta(t) + 3ClO_{4}^{-}(aq) + 24H^{+}(aq)$ + 240 ---4BrO, (aq) + 24H2(aq) + 24g - + 3Cl (aq) + 1214-6900 Burkanica $3ClO_4^-(aq) + 4Br^-(aq) \rightarrow 3Cl^-(aq) + 4BrO_1^-(aq)$ by contiductions: $N_0O_n(g) \rightarrow NO_n^{-1}(aq)$ $N_2O_4(g) \rightarrow 2NO_5^-(mq)$ $N_*O_*(g) + 2H_*O(\ell) \rightarrow 2NO_*^*(aq)$ $N_2O_4(g) + 2H_2O(\ell) \rightarrow 2NO_3^-(aq) + 4H^+(aq)$ $N_2O_\ell(g) + 2H_2O(\ell) + 4OH^*(aq) \rightarrow$ $2NO_3^-(aq) + 4H^+(aq) + 4OH^-(aq)$ $N_2O_4(g) + 2H_2O(\ell) + 4OH^-(aq) \rightarrow$ $2NO_{\tau}^{-}(aq) + 4H_{0}O(\ell)$ $N_1O_\ell(g) + 2H_2O(\ell) + 4OH^-(aq) \rightarrow$ $2NO_3^-(aq) + 4H_2O(\ell) + 2e^$ reduction: $TeO_s^{2-}(aq) \longrightarrow Te(s)$ $TeO_2^{2-}(sq) \longrightarrow Te(s) + 3H_2O(\ell)$ $TeO_1^{2-}(aq) + 6H^*(aq) + 6OH^*(aq) \rightarrow$ $Te(s) + 3H_2O(\ell) + 6OH^{-}(aq)$ $TeO_s^{1-}(aq) + 6H_sO(\ell) \rightarrow$ $Te(s) + 3H_sO(\ell) + 6OH^-(aq)$ $TeO_s^{2-}(aq) + 3H_sO(\ell) \rightarrow Te(s) + 6OH^-(aq)$ $TeO_3^{2-}(aq) + 3H_2O(\ell) + 4e^- \rightarrow Te(s) + 6OH^-(aq)$ multiply oxidation ex. by 2: $2[N_*O_*(g) + 2H_*O(\ell) + 4OH^*(aq) \rightarrow$ $2NO_{\gamma}^{-}(aq) + 4H_{\gamma}O(\ell) + 2e^{-}$ $2N_zO_z(g) + 4H_zO(\ell) + 8OH^-(aq) \rightarrow$ $4NO_3^{-}(aq) + 8H_2O(\ell) + 4e^{-}$ multiply reduction rs. by I: $TeO_s^{2-}(aq) + 6H_sO(\ell) + 4e^- \rightarrow$ $Te(s) + 3H_2O(\ell) + 6OH^-(aq)$ consultations: $2N_2O_4(g) + 4H_2O(\ell) + 8OH^-(aq) + TeO_1^{1-}(aq) +$ 3/H_(O((f) + _6e**----4NO, (aq) + 8H2O(f) + 4e + Te(s) + 6OH (aq) bullance: $2N_2O_4(g) + TeO_3^{2-}(aq) + 2OH^-(aq) \rightarrow$ $Te(s) + 4NO_{s}^{-}(aq) + H_{s}O(\ell)$ c. excidation: $H_2O_2(aq) \longrightarrow O_2(g)$

 $H_2O_2(aq) \longrightarrow O_2(g) + 2H^+(aq)$

 $H.O.(aq) \rightarrow O.(g) + 2H^*(aq) + 2e^-$

reduction: $Cr_*O_2^{-1}(aq) \rightarrow Cr^{3+}(aq)$ $Cr_*O_*^{2-}(aq) \rightarrow 2Cr^{2+}(aq)$ $Cr_*O_*^{2-}(aq) \rightarrow 2Cr^{2+}(aq) + 7H_*O(\ell)$ $Cr_2O_2^{-2}$ (aq) + 14H*(aq) \rightarrow $2Cr^{3+}(aq) + 7H_2O(\ell)$ Cr₂O₂² (aq) + 14H*(aq) + 6e⁻ → $2Cr^{5+}(\omega a) + 7H_2O(0)$ insultiply oxidation ex. by 3: $3[H_2O_2(aq) \rightarrow O_2(g) + 2H^*(aq) + 2e^-]$ $3H_1O_1(aq) \rightarrow 3O_2(g) + 6H^*(aq) + 6e^$ multiply reduction ex. by 1: Cr₂O₂² (aq) + 14H (aq) + 6q - $2Cr^{5+}(aq) + 7H_2O(\ell)$ consistency. 3H₂O₂(aq) + Cr₂O₂² (aq) + 14H*(aq) + 6e²² -- $3O_2(g) + 6H^*(aq) + 6e^- + 2Cr^{3+}(aq) + 7H_2O(\ell)$ balance: $3H_2O_2(\ell) + Cr_2O_2^{2-}(aq) + 8H^+(aq) \rightarrow$ $3O_{r}(g) + 2Cr^{3+}(sg) + 7H_{r}O(f)$ 14. It is a redox reaction because sulfur is both being

Answers to Section 9.3 Review Questions (Student textbook page 617)

marmber of 0 in S(s).

1. Fluorine is the most electronegative element in the periodic table. Therefore, it does not "lose" electrons in chemical bonds. It has seven valence electrons, and thus forms bonds in which it gains one valence electron, satisfying the octet rule. Because it is in the second row of the periodic table, its outer occupied energy level cannot accommodate more than eight electrons (as bromine, iodine, or chlorine can). Therefore, it does not form bonds in which it gains more than one electron, and its oxidation number is always — I.

exidized from an exidation number of -2 in H₂S(g)

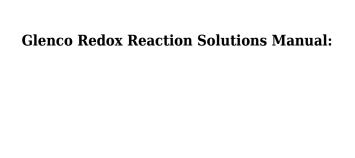
to an oxidation mamber of 0 in S(s) and reduced from

an oxidation number of +4 in SO₂(g) to an oxidation

- 2. a. MacO₂ +4
 - b. +7
 - $c_{-} + 6$
 - dl. $KMmO_n + 7$
- Zero in the diatomic molecule, O₂(g); −1 in peroxides, H₂O₂(aq); −2 in most anions, CO₃^{2−}(aq); +2 in an oxygen difluoride, OF₂(g)
- 4. a. yes
 - the more
 - C. men
 - dl., yees

Glenco Redox Reaction Solutions Manual

J Dewey



Glenco Redox Reaction Solutions Manual: Bestsellers in 2023 The year 2023 has witnessed a noteworthy surge in literary brilliance, with numerous compelling novels captivating the hearts of readers worldwide. Lets delve into the realm of bestselling books, exploring the engaging narratives that have captivated audiences this year. The Must-Read: Colleen Hoovers "It Ends with Us" This poignant tale of love, loss, and resilience has gripped readers with its raw and emotional exploration of domestic abuse. Hoover expertly weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can prevail. Glenco Redox Reaction Solutions Manual: Taylor Jenkins Reids "The Seven Husbands of Evelyn Hugo" This captivating historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reids compelling storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery. Glenco Redox Reaction Solutions Manual: Delia Owens "Where the Crawdads Sing" This mesmerizing coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens weaves a tale of resilience, survival, and the transformative power of nature, captivating readers with its evocative prose and mesmerizing setting. These top-selling novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of captivating stories waiting to be discovered. The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts. The Secret History is a brilliant and thrilling novel that will keep you guessing until the very end. The novel is a warning tale about the dangers of obsession and the power of evil.

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