

Activity A:

Subatomic particles

Get the Gizmo ready:

- Use the arrows to create an atom with two protons, two neutrons, and two electrons.
- Turn on **Show element name**.



Question: What are the properties of protons, neutrons, and electrons?

1. **Observe:** Turn on **Show element symbol** and **Element notation**. Three numbers surround the element symbol: the **mass number** (A), electrical charge (no number is displayed if the atom is neutral), and the **atomic number** (Z).
2. **Investigate:** Watch how the numbers change as you add or remove particles.



- A. Which number is equal to the number of protons in the atom?
- B. How can you calculate the number of neutrons (N) in an atom?
- C. Which particle (proton, neutron, or electron) has a
Positive charge?
Negative charge?

neutrons

by the number of protons

electron

proton

neutron

3. **Explore:** Use the **Gizmo** to create different forms of an element. Each version of an element has the same number of protons, but a different number of neutrons. The element is represented by the element symbol and mass number. For example, Helium-3 has 2 protons and 1 neutron, while Helium-4 has 2 protons and 2 neutrons.

- A. What are the stable isotopes of hydrogen?
- B. What are the stable isotopes of oxygen?
- C. Carbon has three isotopes: carbon-12, carbon-13, and carbon-14.

1 proton and 0 neutrons

8 protons and 8 neutrons

12

4. **Challenge:** Use the **Gizmo** to answer the following questions.

- A. How many electrons are in a neutral atom of Helium?
- B. How many electrons are in an atom of Hydrogen?
- C. What is the mass number of an atom with 3 protons and 3 neutrons?
- D. What is an atom's charge if it contains 11 protons and 10 electrons?
- E. How many electrons are in Hg^{2+} ?

2

1

6

1

10

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M Mark



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