

**Section 5.1: Models of the Atom**

- How many sublevels are in the following principal energy levels? Which sublevels?
  - $n = 1$  **One - s**
  - $n = 2$  **Two - s, p**
  - $n = 3$  **Three - s, p, d**
  - $n = 4$  **Four - s, p, d, f**
  - $n = 5$  **Four - s, p, d, f**
  - $n = 6$  **Three - s, p, d**
  - $n = 7$  **Two - s, p**
- How many orbitals are in the following sublevels?
  - 1s sublevel **One**
  - 2p sublevel **Three**
  - 3d sublevel **Five**
  - 4f sublevel **Seven**
  - 5s sublevel **One**
  - 6p sublevel **Three**
- How many electrons can each of the following sublevels hold (maximum)?
  - s **Two**
  - p **Six**
  - d **Ten**
  - f **Fourteen**
- What is the shape of an s sublevel orbital? **Sphere**
- What is the shape of a p sublevel orbital? **Dumbbell**
- What was the problem with Rutherford's planetary model of the atom? **It could not explain chemical properties**
- Bohr proposed that electrons were found in specific circular paths called **orbits**
- What happens to the distance between energy levels as you increase in energy? **It decreases.**
- What does "electron cloud" mean? **A region of space with the highest probability, or likelihood, of finding an electron. Not an exact path.**
- How do electrons move from energy level to energy level? **Gain or lose energy**

**Section 5.2: Electron Arrangement in Atoms**

- Electrons arrange themselves around the nucleus in order to make the most **stable** arrangement.
- What are the three rules for writing electron configurations? Give the definition of each.
  - Aufbau's Principle** – electrons occupy the lowest energy levels first
  - Pauli Exclusion Principle** – electrons in the same orbital have opposite spin
  - Hund's Rule** – electrons enter orbitals individually before pairing up

# Electrons In Atoms Chapter Test A Worksheet Answers

**RM Cervero**



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**Chapter Resource 2 Chemistry of Life Biology** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2004      *Chemical Misconceptions* Keith Taber, 2002 Part 2 provides strategies for dealing with some of the misconceptions that students have by including ready to use classroom resources

Approaches and Strategies in Next Generation Science Learning Khine, Myint Swe, 2013-01-31 Approaches and Strategies in Next Generation Science Learning examines the challenges involved in the development of modern curriculum models teaching strategies and assessments in science education in order to prepare future students in the 21st century economies This comprehensive collection of research brings together science educators researchers and administrators interested in enhancing the teaching and learning of next generation science      The Science Teacher, 2008 SCC Library has 1964 cur      **Atoms, Molecules & Elements: What Are Atoms? Gr. 5-8** George Graybill, 2015-10-01 This is the chapter slice What Are Atoms from the full lesson plan Atoms Molecules Elements Young scientists will be thrilled to explore the invisible world of atoms molecules and elements Our resource provides ready to use information and activities for remedial students using simplified language and vocabulary Students will label each part of the atom learn what compounds are and explore the patterns in the periodic table of elements to find calcium Ca chlorine Cl and helium He through hands on

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*Atoms and Electrons* John William Navin Sullivan, 1924

**Behavior of Electrons in Atoms** Robin M. Hochstrasser, 1964 Atomic spectra The interactions between atoms and electrons Quantum theory of atomic structure The Pauli principle and the electronic structure of atoms Energy terms and states of atoms Atomic excitation probabilities Collisional processes involving excited atoms The behavior of atoms in magnetic fields Some of the forces between atoms The simplest molecules

*Electrons in Atoms* Karen Horney, 1963

**Electrons in Atoms** G. F. Lothian, 1963

**The Number of Electrons in Atoms** Harold Albert Wilson, 1912

**Electrons and Atoms** John L. Lewis, 1972

*Electrons and the Nucleus* Advanced Physics Project for Independent Learning, 1980 Teacher s guide to accompany one of four second year units of APPIL the Advanced Physics Project for Independent Learning This ILEA project has produced materials which will enable students to study independently for much of their time so that teachers can concentrate on giving individual help where it is needed The materials can be used with any of the A level physics syllabuses except the Nuffield Advanced Physics scheme The teacher s and student s handbooks give information about the whole course The course consists of ten units six for the first year and four for the second year each covering about six weeks work Four units are possible alternative starting points for the course and have preliminary sections for testing students understanding of relevant pre knowledge There are many possible routes through the course which are outlined in the handbooks Each unit has a student s guide and a teacher s guide The student s guides define objectives suggest activities encourage students to use constructively other resources such as textbooks and give instructions for experiments Many

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