

CHAPTER 4 REVIEW

Arrangement of Electrons in Atoms**SECTION 2****SHORT ANSWER** Answer the following questions in the space provided.

1. _____ How many quantum numbers are used to describe the properties of electrons in atomic orbitals?
(a) 1 (c) 3
(b) 2 (d) 4
2. _____ A spherical electron cloud surrounding an atomic nucleus would best represent
(a) an s -orbital. (c) a combination of two different p -orbitals.
(b) a p -orbital. (d) a combination of an s and a p -orbital.
3. _____ How many electrons can an energy level of $n = 4$ hold?
(a) 32 (c) 8
(b) 24 (d) 6
4. _____ How many electrons can an energy level of $n = 2$ hold?
(a) 32 (c) 8
(b) 24 (d) 6
5. _____ Compared with an electron for which $n = 2$, an electron for which $n = 4$ has more
(a) spin. (c) energy.
(b) particle nature. (d) wave nature.
6. _____ According to Bohr, which is the point in the figure below where electrons cannot reside?
(a) point A (c) point C
(b) point B (d) point D



7. _____ According to the quantum theory, point D in the above figure represents:
(a) the fixed position of an electron.
(b) the furthest position from the nucleus that an electron can achieve.
(c) a position where an electron probably exists.
(d) a position where an electron cannot exist.

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