

## 8 • Electron Configurations & Periodicity

### ANSWERS TO STUDY QUESTIONS

- Write the electron configurations of the following elements using the shorthand notation for the noble gas cores.
  - phosphorus [Ne]  $3s^2 3p^3$
  - nickel [Ar]  $3d^8 4s^2$
  - osmium [Xe]  $4f^{14} 5d^6 6s^2$
  - californium [Rn]  $5f^{10} 7s^2$
  - titanium [Ar]  $3d^2 4s^2$
- Which orbital is filled following these orbitals?
  - 3d is followed by 4p
  - 4s is followed by 3d
  - 5p is followed by 6s
  - 5f is followed by 6d
- How many electrons can be accommodated in
  - a d subshell 10; two in each of 5 orbitals
  - a set of f orbitals 14; two in each of 7 orbitals
  - the  $n = 4$  shell 32;  $2s^2$
  - the 7s orbital 2
  - a  $p_x$  orbital? 2
- What is wrong with the following ground state electron configurations?
 

<p>a. <math>\begin{array}{ c c c c } \hline &amp; &amp; &amp; \\ \hline \end{array}</math> <math>\begin{array}{ c } \hline 3d \\ \hline \end{array}</math> <math>\begin{array}{ c c c } \hline &amp; &amp; \\ \hline \end{array}</math> <math>\begin{array}{ c } \hline 4s \\ \hline \end{array}</math></p> <p>b. <math>\begin{array}{ c c c c c } \hline 1\downarrow &amp; 1\downarrow &amp; 1\downarrow &amp; 1\downarrow &amp; 1\downarrow \\ \hline \end{array}</math> <math>\begin{array}{ c } \hline 3d \\ \hline \end{array}</math> <math>\begin{array}{ c } \hline 4s \\ \hline \end{array}</math> <math>\begin{array}{ c c c } \hline 1\downarrow &amp; 1\downarrow &amp; 1\downarrow \\ \hline \end{array}</math> <math>\begin{array}{ c } \hline 4p \\ \hline \end{array}</math></p> <p>c. <math>\begin{array}{ c c c c c } \hline 1\downarrow &amp; 1\downarrow &amp; 1\downarrow &amp; 1\downarrow &amp; 1\downarrow \\ \hline \end{array}</math> <math>\begin{array}{ c } \hline 3d \\ \hline \end{array}</math> <math>\begin{array}{ c } \hline 4s \\ \hline \end{array}</math> <math>\begin{array}{ c c c } \hline 1\downarrow &amp; 1\downarrow &amp; 1\downarrow \\ \hline \end{array}</math> <math>\begin{array}{ c } \hline 4p \\ \hline \end{array}</math></p> <p>d. <math>\begin{array}{ c c c c c } \hline 1\downarrow &amp; 1\downarrow &amp; 1\downarrow &amp; 1\downarrow &amp; 1\downarrow \\ \hline \end{array}</math> <math>\begin{array}{ c } \hline 3d \\ \hline \end{array}</math> <math>\begin{array}{ c } \hline 4s \\ \hline \end{array}</math> <math>\begin{array}{ c c c } \hline &amp; &amp; \\ \hline \end{array}</math> <math>\begin{array}{ c } \hline 4p \\ \hline \end{array}</math></p> <p>e. <math>\begin{array}{ c c c c c } \hline 1\downarrow &amp; 1\downarrow &amp; 1\downarrow &amp; 1\downarrow &amp; 1\downarrow \\ \hline \end{array}</math> <math>\begin{array}{ c } \hline 3d \\ \hline \end{array}</math> <math>\begin{array}{ c } \hline 4s \\ \hline \end{array}</math> <math>\begin{array}{ c c c } \hline 1\downarrow &amp; &amp; \\ \hline \end{array}</math> <math>\begin{array}{ c } \hline 4p \\ \hline \end{array}</math></p>	<p>a. OK, 4s is filled before 3d.</p> <p>b. 4s should be filled before 4p subshell.</p> <p>c. 4s cannot have 2 electrons with same spin.</p> <p>d. OK, stability of half-filled 3d subshell.</p> <p>e. 4p should first fill one orbital at a time.</p>
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- How many unpaired electrons are there in
  - a nitrogen atom 3 unpaired electrons ( $2s^2 2p^3$ )
  - an iodine atom 1 unpaired electrons ( $5s^2 5p^5$ )
  - a nickel (II) cation 2 unpaired electrons ( $3d^8$ )
  - an oxide ion no unpaired electrons (closed shell [Ne])

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