

## Section 8.5

### Objectives

- **Describe** how electronegativity is used to determine bond type.
- **Compare and contrast** polar and nonpolar covalent bonds and polar and nonpolar molecules.
- **Generalize** about the characteristics of covalently bonded compounds.

## Review Vocabulary

**electronegativity:** the relative ability of an atom to attract electrons in a chemical bond

### New Vocabulary

polar covalent bond

## Electronegativity and Polarity

**MAIN Idea** A chemical bond's character is related to each atom's attraction for the electrons in the bond.

**Real-World Reading Link** The stronger you are, the more easily you can do pull-ups. Just as people have different abilities for doing pull-ups, atoms in chemical bonds have different abilities to attract (pull) electrons.

## Electron Affinity, Electronegativity, and Bond Character

The type of bond formed during a reaction is related to each atom's attraction for electrons. Electron affinity is a measure of the tendency of an atom to accept an electron. Excluding noble gases, electron affinity increases with increasing atomic number within a period and decreases with increasing atomic number within a group. The scale of electronegativities—shown in **Figure 8.20**—allows chemists to evaluate the electron affinity of specific atoms in a compound. Recall from Chapter 6 that electronegativity indicates the relative ability of an atom to attract electrons in a chemical bond. Note that electronegativity values were assigned, whereas electron affinity values were measured.

**Electronegativity** The version of the periodic table of the elements shown in **Figure 8.20** lists electronegativity values. Note that fluorine has the greatest electronegativity value (3.98), while francium has the least (0.7). Because noble gases do not generally form compounds, individual electronegativity values for helium, neon, and argon are not listed. However, larger noble gases, such as xenon, sometimes bond with highly electronegative atoms, such as fluorine.

**Figure 8.20** Electronegativity values are derived by comparing an atom's attraction for shared electrons to that of a fluorine's atom attraction for shared electrons. Note that the electronegativity values for the lanthanide and actinide series, which are not shown, range from 1.12 to 1.7.

1 <b>H</b> 2.20																	5 <b>B</b> 2.04	6 <b>C</b> 2.55	7 <b>N</b> 3.04	8 <b>O</b> 3.44	9 <b>F</b> 3.98											
3 <b>Li</b> 0.98	4 <b>Be</b> 1.57															13 <b>Al</b> 1.61	14 <b>Si</b> 1.90	15 <b>P</b> 2.19	16 <b>S</b> 2.58	17 <b>Cl</b> 3.16												
11 <b>Na</b> 0.93	12 <b>Mg</b> 1.31															19 <b>K</b> 0.82	20 <b>Ca</b> 1.00	21 <b>Sc</b> 1.36	22 <b>Ti</b> 1.54	23 <b>V</b> 1.63	24 <b>Cr</b> 1.66	25 <b>Mn</b> 1.55	26 <b>Fe</b> 1.83	27 <b>Co</b> 1.88	28 <b>Ni</b> 1.91	29 <b>Cu</b> 1.90	30 <b>Zn</b> 1.65	31 <b>Ga</b> 1.87	32 <b>Ge</b> 2.05	33 <b>As</b> 2.18	34 <b>Se</b> 2.55	35 <b>Br</b> 2.96
39 <b>Rb</b> 0.82	40 <b>Sr</b> 0.95	39 <b>Y</b> 1.27	40 <b>Zr</b> 1.33	41 <b>Nb</b> 1.6	42 <b>Mo</b> 2.16	43 <b>Tc</b> 2.10	44 <b>Ru</b> 2.2	45 <b>Rh</b> 2.28	46 <b>Pd</b> 2.20	47 <b>Ag</b> 1.93	48 <b>Cd</b> 1.69	49 <b>In</b> 1.78	50 <b>Sn</b> 1.96	51 <b>Sb</b> 2.05	52 <b>Te</b> 2.1	53 <b>I</b> 2.66																
55 <b>Cs</b> 0.79	56 <b>Ba</b> 0.89	57 <b>La</b> 1.10	72 <b>Hf</b> 1.3	73 <b>Ta</b> 1.5	74 <b>W</b> 1.7	75 <b>Re</b> 1.9	76 <b>Os</b> 2.2	77 <b>Ir</b> 2.2	78 <b>Pt</b> 2.2	79 <b>Au</b> 2.4	80 <b>Hg</b> 1.9	83 <b>Tl</b> 1.8	84 <b>Pb</b> 1.8	85 <b>Bi</b> 1.9	86 <b>Po</b> 2.0	87 <b>At</b> 2.2																
87 <b>Fr</b> 0.7	88 <b>Ra</b> 0.9	89 <b>Ac</b> 1.1																														

# Electronegativity Polarity Study Guide Answers

**B Lingard**



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## **Table of Contents Electronegativity Polarity Study Guide Answers**

1. Understanding the eBook Electronegativity Polarity Study Guide Answers
  - The Rise of Digital Reading Electronegativity Polarity Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Electronegativity Polarity Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Electronegativity Polarity Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Electronegativity Polarity Study Guide Answers
  - Personalized Recommendations
  - Electronegativity Polarity Study Guide Answers User Reviews and Ratings
  - Electronegativity Polarity Study Guide Answers and Bestseller Lists
5. Accessing Electronegativity Polarity Study Guide Answers Free and Paid eBooks
  - Electronegativity Polarity Study Guide Answers Public Domain eBooks
  - Electronegativity Polarity Study Guide Answers eBook Subscription Services
  - Electronegativity Polarity Study Guide Answers Budget-Friendly Options

6. Navigating Electronegativity Polarity Study Guide Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Electronegativity Polarity Study Guide Answers Compatibility with Devices
  - Electronegativity Polarity Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Electronegativity Polarity Study Guide Answers
  - Highlighting and Note-Taking Electronegativity Polarity Study Guide Answers
  - Interactive Elements Electronegativity Polarity Study Guide Answers
8. Staying Engaged with Electronegativity Polarity Study Guide Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Electronegativity Polarity Study Guide Answers
9. Balancing eBooks and Physical Books Electronegativity Polarity Study Guide Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Electronegativity Polarity Study Guide Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Electronegativity Polarity Study Guide Answers
  - Setting Reading Goals Electronegativity Polarity Study Guide Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Electronegativity Polarity Study Guide Answers
  - Fact-Checking eBook Content of Electronegativity Polarity Study Guide Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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