# Electron Configurations - KEY

Note: The electron configurations in this worksheet assume that lanthanum (La) is the first element in the 4f block and that actinium (Ac) is the first element in the 5f block. If your periodic table doesn't agree with this, your answers for elements near the f-orbitals may be slightly different (one more/less electron for the f-orbital).

1s22s22p63s1 13 sodium. 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>6</sup>4s<sup>2</sup>3d<sup>6</sup> 2) iron: 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>6</sup>4s<sup>2</sup>3d<sup>10</sup>4p<sup>5</sup> bromine. 3) 1s22s22p63s23p64s23d104p65s24d105p66s2 4) barium. 5) neptunium 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>6</sup>4s<sup>2</sup>3d<sup>10</sup>4p<sup>6</sup>5s<sup>2</sup>4d<sup>10</sup>5p<sup>6</sup>6s<sup>2</sup>4f<sup>14</sup>5d<sup>10</sup>6p<sup>6</sup>7s<sup>2</sup>5f<sup>5</sup> also might be written (due to the close energy of d and I and similarities of elements versus predictable memorized config. 1s22s2p63s23p64s23d104p65s24d105p66s24f145d106p67s26d15f4 [Arl 4s<sup>2</sup>3d<sup>2</sup> 60) cobalt [Krl 5s24d9 70 silluer [Kr] 5s24d105p4 80 tellurium. [Rn] 7s2 90 radium lawrencium [Rn] 7s<sup>2</sup>5f<sup>14</sup>6d<sup>1</sup> ht be written (due to the close energy of d and f and similarities of elements versus predictable memorized config. [Rn] 7s26d15f14 1s22s2p63s23p4 sulfur 11) 1s22s2p63s23p64s23d104p65s1 rubidium 12) [Kr] 5s24d105p3 antimony 13) [Xe] 6s24f145d6 osmium 14) [Rn] 7s25f11 einsteinium note: no d'a so the p. table referenced has Ac as 5f1 15% 1s22s2p63s23p64s24d104p5 not valid (take a look at "4d" must be 3 d not 4th 160 1s22s2p53s3d5 not valid (3p comes after 3s); also no 3 electrons in S 17) [Ra] 7s25f8 not valid (radium isn't a noble gas) 18) [Kr] 5s24d105p5 valid 19) 20) [Xe] not valid (an element can't be its own electron configuration)

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Robert C. Smoot, Smoot, Richard G. Smith, Jack Price

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