Name: \_\_\_\_\_ Date: \_\_\_\_\_

Electron configuration worksheet with answer key

Write the full electron configurations for each element.

Potassium (K)

1s 2s 2p 3s 3p 4s

· Sulfur (S)

1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>4</sup>

Fluorine (F)

1s<sup>2</sup>2s<sup>2</sup>2p<sup>5</sup>

Silicon (Si)

1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>2</sup>

Write the shorthand (noble gas) electron configurations.

Calcium (Ca)

[Ar]4s2

Iron (Fe)

[Ar]4s<sup>2</sup>3d<sup>6</sup>

Nickel (Ni)

[Ar]4s<sup>2</sup>3d<sup>8</sup>

Copper (Cu)

[Ar]4s 3d 10

Identify elements based on their electron configurations.

- 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup> 3p<sup>6</sup> 4s<sup>1</sup> K (Potassium)
- [Ne] 3s<sup>2</sup> 3p<sup>3</sup>

P (Phosphorus)

[Ar] 4s<sup>2</sup> 3d<sup>8</sup>

(Nickel)

[Kr] 5s¹

(Rubidium)

# **Electron Configuration Lab Answers**

**Richard H. Pantell** 

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