

Matter, Chemical Trends and Chemical Bonding Unit Test

Knowledge	Thinking	Application	Communication
11/14	12/17	13/16	12/12

Thinking/Investigation (17 marks)

1. An element is composed of 70% of an isotope with an atomic mass of 69.0u. If the other isotope has 2 more neutrons per atom, what is the average atomic mass of the element (4 marks)?

$$AAM = (M_1 \times \text{Abundance}) + (M_2 \times \text{Abundance})$$

$$AAM = (69.0u \times 70\%) + (71.0u \times 30\%)$$

$$AAM = 69.6u$$

4

2. Fill in the following chart (13 marks)

Formula	Lewis Structure	Name(s) of Shape	Overall Polarity of Molecule
SBr_2	$\text{:}\ddot{\text{Br}}\text{--}\ddot{\text{S}}\text{--}\ddot{\text{Br}}\text{:}$ $\text{:}\ddot{\text{Br}}\text{:}$	Trigonal planar ✓	Non-Polar ✓
PO_3^{3-}	$\left[\text{:}\ddot{\text{O}}\text{--}\ddot{\text{P}}\text{--}\ddot{\text{O}}\text{:} \right]^{-3}$ $\text{:}\ddot{\text{O}}\text{:}$	Trigonal pyramidal ✓	Non-Polar ✗
HCP	$\text{H--}\ddot{\text{C}}\equiv\text{P:}$ why so many ✓	Bent ✗	Non-Polar ✓
CH_3NH_2	$\begin{array}{c} \text{H} & \text{H} \\ & \\ \text{H--C--N:} \\ & \\ \text{H} & \text{H} \end{array}$ ✓	Trigonal planar ✓	Non-Polar ✗

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Table of Contents Grade 11 Chemistry Unit 1 Practice Test

1. Understanding the eBook Grade 11 Chemistry Unit 1 Practice Test
 - The Rise of Digital Reading Grade 11 Chemistry Unit 1 Practice Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Grade 11 Chemistry Unit 1 Practice Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grade 11 Chemistry Unit 1 Practice Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Grade 11 Chemistry Unit 1 Practice Test
 - Personalized Recommendations
 - Grade 11 Chemistry Unit 1 Practice Test User Reviews and Ratings
 - Grade 11 Chemistry Unit 1 Practice Test and Bestseller Lists
5. Accessing Grade 11 Chemistry Unit 1 Practice Test Free and Paid eBooks

- Grade 11 Chemistry Unit 1 Practice Test Public Domain eBooks
 - Grade 11 Chemistry Unit 1 Practice Test eBook Subscription Services
 - Grade 11 Chemistry Unit 1 Practice Test Budget-Friendly Options
6. Navigating Grade 11 Chemistry Unit 1 Practice Test eBook Formats
- ePub, PDF, MOBI, and More
 - Grade 11 Chemistry Unit 1 Practice Test Compatibility with Devices
 - Grade 11 Chemistry Unit 1 Practice Test Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Grade 11 Chemistry Unit 1 Practice Test
 - Highlighting and Note-Taking Grade 11 Chemistry Unit 1 Practice Test
 - Interactive Elements Grade 11 Chemistry Unit 1 Practice Test
8. Staying Engaged with Grade 11 Chemistry Unit 1 Practice Test
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grade 11 Chemistry Unit 1 Practice Test
9. Balancing eBooks and Physical Books Grade 11 Chemistry Unit 1 Practice Test
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Grade 11 Chemistry Unit 1 Practice Test
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Grade 11 Chemistry Unit 1 Practice Test
- Setting Reading Goals Grade 11 Chemistry Unit 1 Practice Test
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Grade 11 Chemistry Unit 1 Practice Test
- Fact-Checking eBook Content of Grade 11 Chemistry Unit 1 Practice Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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