

Grade 11 Business Studies Past Papers And Memos

Jeffery Wama



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buffers typically consist of an acid base pair with the acid and base differing by the presence or absence of a proton a conjugate acid base pair for instance one of the buffers that maintain the ph of human blood involves carbonic acid H_2CO_3 and its conjugate base the bicarbonate ion HCO_3^-

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$$\text{acid} + \text{base} \rightarrow \text{water} + \text{salt}$$

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