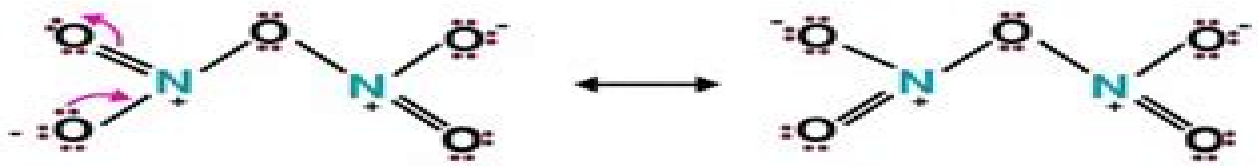

55. Phosphorus has three allotropic forms — (i) white phosphorus (ii) red phosphorus and (iii) black phosphorus. Write the difference between white and red phosphorus on the basis of their structure and reactivity.
Solution:
1. White phosphorus
(i) It is soft waxy solid having garlic order
(ii) It is poisonous
(iii) Has low melting point and boiling point because P4 molecules are held together by weak van der Waal forces of attraction

2. Red Phosphorus
(i) It is hard, crystalline, odourless solid
(ii) It is non-poisonous
(iii) It has a high melting point because of its polymeric structure.

56. Give an example to show the effect of concentration of nitric acid on the formation of oxidation product.
Solution:
$4Zn + 10HNO_3(dil) \rightarrow 4Zn(NO_3)_2 + NH_4NO_3 + 3H_2O$

$4Zn + 10HNO_3(dil) \rightarrow 4Zn(NO_3)_2 + N_2O + 5H_2O$

$4Zn + 4HNO_3(conc.) \rightarrow Zn(NO_3)_2 + NO_2 + 2H_2O$

57. PCl_5 reacts with finely divided silver on heating and a white silver salt is obtained, which dissolves on adding excess aqueous NH_3 solution. Write the reactions involved to explain what happens.
Solution:
PCl_5 reacts with silver to form white silver salt(AgCl). Which than dissolves on adding excess aqueous NH_3 to form a soluble complex
$PCl_5 + 2Ag \rightarrow 2AgCl(white\ ppt.) + PCl_3$
$AgCl + 2NH_3(aq) \rightarrow [Ag(NH_3)]^+Cl^-$

58. Phosphorus forms a number of oxoacids. Out of these oxoacidsphosphinicacid has strong

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