

199) A student finds the ΔG of hydrolysis of ATP is -30.5 kJ/mol at pH 7.0. Is this value the same as the ΔG of phosphorylation? At which condition would phosphorylation be most favorable?

(a) Indeed does the same phosphorylation report

_____ ΔG for H_2O is -30.5 kJ/mol _____ ΔG for ATP _____ ΔG for $ADP + P_i$ _____ ΔG for $ATP \rightarrow ADP + P_i$ _____ ΔG

(b) What is the relative free energy of ATP and ADP at pH 7.0?

_____ ATP is 10^4 times more than ADP _____ ΔG

(c) Calculate the phosphorylation of ATP from the following data:

$$\frac{1.2 \times 10^{-3}}{2 \times 10^{-3}} = \frac{[ATP]}{[ADP + P_i]} \quad \Delta G = -30.5 \text{ kJ/mol}$$

_____ ΔG for $ATP \rightarrow ADP + P_i$ is -30.5 kJ/mol _____ ΔG for $ADP + P_i \rightarrow ATP$ is 30.5 kJ/mol _____ ΔG for $ATP \rightarrow ADP + P_i$ is -30.5 kJ/mol _____ ΔG

(d) What is the free energy of ATP at pH 7.0? Is this value the same as the ΔG of phosphorylation?

(a) Phosphorylation is the same as ΔG of ATP at pH 7.0

_____ ΔG for $ATP \rightarrow ADP + P_i$ is -30.5 kJ/mol _____ ΔG for $ADP + P_i \rightarrow ATP$ is 30.5 kJ/mol _____ ΔG for $ATP \rightarrow ADP + P_i$ is -30.5 kJ/mol _____ ΔG

(b) Phosphorylation is the same as ΔG of ATP at pH 7.0

_____ ΔG for $ATP \rightarrow ADP + P_i$ is -30.5 kJ/mol _____ ΔG for $ADP + P_i \rightarrow ATP$ is 30.5 kJ/mol _____ ΔG for $ATP \rightarrow ADP + P_i$ is -30.5 kJ/mol _____ ΔG

_____ ΔG for $ATP \rightarrow ADP + P_i$ is -30.5 kJ/mol _____ ΔG for $ADP + P_i \rightarrow ATP$ is 30.5 kJ/mol _____ ΔG for $ATP \rightarrow ADP + P_i$ is -30.5 kJ/mol _____ ΔG

_____ ΔG for $ATP \rightarrow ADP + P_i$ is -30.5 kJ/mol _____ ΔG for $ADP + P_i \rightarrow ATP$ is 30.5 kJ/mol _____ ΔG for $ATP \rightarrow ADP + P_i$ is -30.5 kJ/mol _____ ΔG

(Total: 100)



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