

## GROUP 2 ELEMENTS AND THEIR COMPOUNDS

- show their outermost electron placed on a subshell
- readily react to lose 2 electrons
  - $\text{M} \rightarrow \text{M}^{2+} + 2\text{e}^{-}$

### Trends down the group:

- $1^{\text{st}}$  ionisation energy decreases
  - Distance between outermost electron and nucleus increases
  - More shielding
  - Counteractive effect of increasing nuclear charge
- Atomic radius increases
- Melting point generally decreases
  - Much lower than all transition metals

### Physical properties:

- Good conductors of heat and electricity
  - Sea of delocalised electrons
- Form white compounds
  - Transition metals form coloured compounds
- Oxidation number of +2
- All compounds are ionic or almost ionic
  - Be compounds show more covalent character because small highly charged cation polarises anion shell of anion
- Oxides / hydroxides are basic
  - React with acids (proton acceptors)
  - Example:
    - $\text{BaO}$
    - $\text{BaO} + 2\text{H}^{+} \rightarrow \text{Ba}^{2+} + \text{H}_2\text{O}$  which is exothermic
- Density =  $2.83 \text{ g/cm}^3$ 
  - Average density of transition metals is about  $7 \text{ g/cm}^3$
- Compared to group 2 metals:
  - Harder + denser
  - Higher melting point
    - Metallic bond is stronger (higher charge on cation & more electrons delocalised)

### Chemical properties:

- Strong reducing agents
- Oxidation is losing electrons
- Reduction is gaining electrons:
  - Reducing agent - donates electrons (reducing other substances)
  - Oxidising agent - accepts electrons (oxidising itself)
- E.g.
  - $\text{ZnO} + \text{CO} \rightarrow \text{Zn} + \text{CO}_2$ 
    - $\text{ZnO} \rightarrow \text{Zn}^{2+} + 2\text{O}^{2-}$
    - $\text{CO} + 2\text{O}^{2-} \rightarrow \text{CO}_2^{2-}$

# F321 Ocr A Chemistry May 2014 Paper

**Didier Musso**



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