Name:	Class:	Date:

#### Chapter 5 Chemistry: Electrons in Atoms

#### Matching.

Match each item with the correct statement below.

- Pauli exclusion principle

- electron configuration
- aufbau principle ground state d.
- 1. each orbital has at most two electrons
- lowest energy level.
- 3. tendency of electrons to enter orbitals of lowest energy first
  - 4. arrangement of electrons around atomic nucleus

#### Multiple Choice

Identify the choice that best completes the statement or answers the question.

- What is the maximum number of electrons that can occupy the n = 3 shell?
  - b.

c. 8

18

- 6. Which of the following elements is a d-block element?
  - Cd

b.

Cs

CI

- Which element has the electron configuration Is<sup>2</sup>2s<sup>3</sup>2p<sup>6</sup>3d<sup>66</sup>4s<sup>2</sup>?
  - Ag
- b. Cd.

- Circ
- 8. Which element has the following ground state electron configuration?

Co

48.8 Cu

Za

- shell is the lowest that may contain d-orbitals.

Fig.

b. 4

- c. 3
- 10. In Bohr's model of the atom, where are the electrons and protons located?
  - The electrons move around the protons, which are at the center of the atom.
  - The electrons and protons move throughout the atom.
  - The electrons occupy fixed positions around the protons, which are at the center of the
  - The electrons and protons are located throughout the atom, but they are not free to move.
- 11. How many energy sublevels are in the second principal energy level?

h: 2

- d. 4
- 12. What is the maximum number of orbitals in the p sublevel?
  - ъ. 3

c. 4 d. 5

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**C** Cleary

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