8 • Electron Configurations & Periodicity

ANSWERS TO STUDY QUESTIONS

 Write the electron configurations of the following elements using the shorthand notation for the noble gas cores.

a. phosphorus [Ne] 3s2 3p3

b. nickel [Ar] 3d⁸ 4s²

c. osmium [Xe] 4f⁴⁴ 5d⁶ 6s²

- d. californium [Rn] 5f¹⁰ 7s² e. titanium [Ar] 3d² 4s²
- 2. Which orbital is filled following these orbitals?
- a. 3d is followed by 4p
 b. 4s is followed by 3d

c. 5p is followed by 6s

- d. 5f is followed by 6d
- How many electrons can be accommodated in a. a d subshell 10; two in each of 5 orbitals
- b. a set of f orbitals 14; two in each of 7 orbitals c. the = 4 shell 32; 2,2
- d. the 7s orbital 2

e. a p., orbital? 2

4. What is wrong with the following ground state electron configurations?

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a. OK, 4s is filled before 3d.

b. 4s should be filled before 4p subshell.

c. 4s cannot have 2 electrons with same spin.

d. OK, stability of half-filled 3d subshell.

e. 4p should first fill one orbital at a time.

- 5. How many unpaired electrons are there in
 - a. a nitrogen atom 3 unpaired electrons (2s2 2p3)
- b. an iodine atom 1 unpaired electrons (5s² 5p⁵)
 - c. a nickel (II) cation 2 unpaired electrons (3d*)
- d. an oxide ion no unpaired electrons (closed shell [Ne])

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Leo J. Malone, Theodore O. Dolter

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