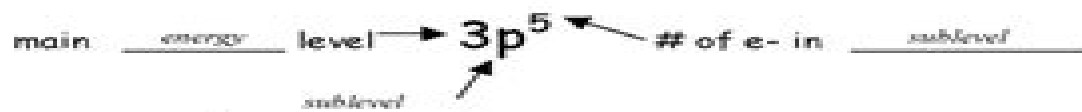


- There are four types of orbitals:  
**s** : shaped like a sphere  
An E level can contain only 1 s orbital, making up the "s sublevel".  
**p** : shaped like dumbbells  
An E level can contain 3 p orbitals, making up the "p sublevel".  
**d** : shaped like double dumbbells  
An E level can contain 5 d orbitals, making up the "d sublevel".  
**f** : too complex to draw or describe  
An E level can contain 7 f orbitals, making up the "f sublevel".
- Each orbital can hold a maximum of 2 electrons. Since both electrons have a negative charge, they repel. What keeps them from flying apart?  
*Each electron spins on its axis. One spins clockwise and the other spins counter-clockwise. When charged particles spin, they act like tiny magnets. Since the two electrons spin in opposite directions, one acts like the north pole of a magnet and the other acts like the south pole. This makes the electrons attract.*
- Since each orbital can hold 2 electrons:  
The "s sublevel" can hold 2 electrons.  
The "p" sublevel can hold 6 electrons.  
The "d" sublevel can hold 10 electrons.  
The "f" sublevel can hold 14 electrons.

We use this notation to describe an electron:



How are electrons distributed within a sublevel?

According to Hund's Rule, each orbital within a sublevel is half-filled before any is filled.

# Electron Distributions Review Answer Key

**Xiaolong Qi**



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