

Name KEY

Orbital Diagrams

Orbital diagrams show the electronic configuration and electron pairings within an atom. Each box represents an orbital. Each orbital can only hold two electrons. Electrons are represented by arrows, which can point up (spin up) or point down (spin down). Remember the two rules for orbital diagrams: *Pauli Exclusion Principle* - no two electrons with the same spin can occupy the same orbital, and *Hund's Rule* - every orbital in the same sublevel must have one electron before electrons will pair.

First label the boxes with their principal energy level number and sublevel letter. Then start to fill in the electrons (put arrows in the boxes). When your orbital diagram is finished, write the final electron configuration on the line below the orbital diagram.

Al

↑↓	↑↓	↑↓	↑↓	↑↓	↑		
1s	2s	2p			3s		

 Electron configuration: $1s^2 2s^2 2p^6 3s^2 3p^1$

C

↑↓	↑↓	↑	↑				
1s	2s	2p					

 Electron configuration: $1s^2 2s^2 2p^2$

Si

↑↓	↑↓	↑↓	↑↓	↑↓	↑	↑	
1s	2s	2p			3s		

 Electron configuration: $1s^2 2s^2 2p^6 3s^2 3p^2$

Cl

↑↓	↑↓	↑↓	↑↓	↑↓	↑↓	↑↓	↑
1s	2s	2p			3s		

 Electron configuration: $1s^2 2s^2 2p^6 3s^2 3p^5$

O

↑↓	↑↓	↑↓	↑	↑			
1s	2s	2p					

 Electron configuration: $1s^2 2s^2 2p^4$

Draw an orbital diagram for the following atoms on your own.

B

↑↓	↑↓	↑		
1s	2s	2p		

 Electron configuration: $1s^2 2s^2 2p^1$

K

↑↓	↑↓	↑↓	↑↓	↑↓	↑↓	↑↓	↑	
1s	2s	2p			3s	3p		4s

 Electron configuration: $1s^2 2s^2 2p^6 3s^2 3p^4 4s^1$

Electron Configuration Elements Atoms And Ions

Answer Key

Hongru Du



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