

1. *Journal of the American Medical Association*, 1997; 277: 1039-1043.

- [illegible]

[illegible]

3. Which type of bonding interaction is exemplified by the H_2O^+ ion?
 - a. Covalent
 - b. Ionic
 - c. Van der Waals
4. Which type of bonding interaction is exemplified by the hydrogen bonding between two water molecules?
 - a. Covalent
 - b. Ionic
 - c. Van der Waals
5. Which interaction is the most important in stabilizing the structure of DNA?
 - a. Hydrogen bonding
 - b. Ionic bonding
 - c. Covalent bonding
6. Which of the following elements is most likely to form a covalent bond?
 - a. Sodium
 - b. Chlorine
 - c. Calcium
7. Which of the following elements is most likely to form an ionic bond?
 - a. Sodium
 - b. Chlorine
 - c. Calcium
8. Which of the following elements is most likely to form a metallic bond?
 - a. Sodium
 - b. Chlorine
 - c. Calcium
9. Which of the following elements is most likely to form a network solid?
 - a. Carbon
 - b. Silicon
 - c. Boron
10. Which of the following elements is most likely to form a molecular solid?
 - a. Oxygen
 - b. Nitrogen
 - c. Fluorine
11. Which of the following elements is most likely to form a covalent network solid?
 - a. Silicon
 - b. Boron
 - c. Carbon
12. Which of the following elements is most likely to form a metallic solid?
 - a. Sodium
 - b. Calcium
 - c. Magnesium
13. Which of the following elements is most likely to form a molecular solid?
 - a. Oxygen
 - b. Nitrogen
 - c. Fluorine
14. Which of the following elements is most likely to form a covalent network solid?
 - a. Silicon
 - b. Boron
 - c. Carbon
15. Which of the following elements is most likely to form a metallic solid?
 - a. Sodium
 - b. Calcium
 - c. Magnesium
16. Which of the following elements is most likely to form a molecular solid?
 - a. Oxygen
 - b. Nitrogen
 - c. Fluorine
17. Which of the following elements is most likely to form a covalent network solid?
 - a. Silicon
 - b. Boron
 - c. Carbon
18. Which of the following elements is most likely to form a metallic solid?
 - a. Sodium
 - b. Calcium
 - c. Magnesium
19. Which of the following elements is most likely to form a molecular solid?
 - a. Oxygen
 - b. Nitrogen
 - c. Fluorine
20. Which of the following elements is most likely to form a covalent network solid?
 - a. Silicon
 - b. Boron
 - c. Carbon

1. ☐ **Yes**
 2. ☐ **No**
 3. ☐ **Other**

- g) Sketch a typical bonding scheme in a conjugated polycyclic aromatic hydrocarbon. Briefly describe the electronic structure of the bonding system.
 - i) π -conjugation
 - ii) delocalization
 - iii) Hückel
- h) Characterize aromaticity in terms of aromaticity requirements:
 - i) Hückel's 4n+2 rule
 - ii) aromaticity in terms of delocalization
 - iii) aromaticity in terms of resonance
- i) Give a definition of aromaticity and describe the aromaticity of benzene. Discuss the aromaticity of furan, pyrrole, and thiophene.
 - i) π -conjugation
 - ii) delocalization
 - iii) Hückel
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 - i) π -conjugation
 - ii) delocalization
 - iii) Hückel

Figure 1

4. Calculate work function ϕ in eV (rounded to 2 significant figures):
a. $\lambda_{\text{max}} = 1.9 \times 10^{-7} \text{ m}$
5. Arrange the following elements in order of decreasing electronegativity:
arsenic, iron, boron, nitrogen, calcium
6. One molecule in a polyatomic ionic compound has a charge that is other than ± 1 . The electronegativity values for the elements which are in the following compound is a good reason:
a. Ca^{2+}
b. Fe^{3+}
c. N^{3-}
7. One electronegativity value is determined with reference to following elements (not shown):
argon, chlorine, bromine, iodine
a. Fe
b. Ca
c. F
8. Electronegativity values of bonding elements and the number of nonbonding electrons in the structure of CH_4 .
9. Electronegativity values of bonding elements and the number of nonbonding electrons in the structure of H_2O .
10. For a molecular substance with the atoms arranged as follows, draw the Lewis symbol. The placement of bonded is unimportant in the order of giving the nonbonding pairs. Award a maximum of 2 if the is electronegative change through whole atoms.
 $\text{H}-\text{O}-\text{N}-\text{H}$
11. For possible electron-dot structures, are shown for the compound $\text{Fe}_2\text{Br}_2\text{Cl}_2$. Explain why one structure shows two bridging bonds/atoms and

Figure 1

1. Draw the following structures for each of the following polymers (approximately 100 Å long). Identify the repeating structure.
 - a. PE
 - b. PPS
 - c. PVC
 - d. PMMA
 - e. PTFE
2. Draw the following polymers (approximately 100 Å long). Indicate the repeating structure (and draw it) and the monomer.
 - a. PE
 - b. PPS
 - c. PVC
 - d. PMMA
 - e. PTFE
3. Draw the repeating structure and the monomer of each of the following polymers.
 - a. PE
 - b. PPS
 - c. PVC
 - d. PMMA
 - e. PTFE
4. Draw the repeating structure and the monomer of each of the following polymers.
 - a. PE
 - b. PPS
 - c. PVC
 - d. PMMA
 - e. PTFE
5. Draw the repeating structure and the monomer of each of the following polymers.
 - a. PE
 - b. PPS
 - c. PVC
 - d. PMMA
 - e. PTFE
6. Draw the repeating structure and the monomer of each of the following polymers.
 - a. PE
 - b. PPS
 - c. PVC
 - d. PMMA
 - e. PTFE
7. Draw the repeating structure and the monomer of each of the following polymers.
 - a. PE
 - b. PPS
 - c. PVC
 - d. PMMA
 - e. PTFE
8. Draw the repeating structure and the monomer of each of the following polymers.
 - a. PE
 - b. PPS
 - c. PVC
 - d. PMMA
 - e. PTFE
9. Draw the repeating structure and the monomer of each of the following polymers.
 - a. PE
 - b. PPS
 - c. PVC
 - d. PMMA
 - e. PTFE
10. Draw the repeating structure and the monomer of each of the following polymers.
 - a. PE
 - b. PPS
 - c. PVC
 - d. PMMA
 - e. PTFE

- (a) The following is a probability distribution for the number of children in a family. Verify that the distribution is a probability distribution. If it is not, explain why.

x	0	1	2	3	4
$P(x)$	0.15	0.25	0.35	0.15	0.10
- (b) The following is a probability distribution for the number of children in a family. Verify that the distribution is a probability distribution. If it is not, explain why.

x	0	1	2	3	4
$P(x)$	0.15	0.25	0.35	0.15	0.10
- (c) Determine the mean of the following distribution and the standard deviation of the following distribution for the structure of 2000.
- (d) Determine the mean of the following distribution and the standard deviation of the following distribution for the structure of 2000.
- (e) For a continuous random variable, the mean is 10 and the standard deviation is 2. Find the probability that the value of the random variable is greater than 12.
- (f) For a continuous random variable, the mean is 10 and the standard deviation is 2. Find the probability that the value of the random variable is less than 12.

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- (1) How does the geometry of electron distribution in the π -conjugated backbone influence the electronic transport properties? Justify your answer.
- (2) How does the geometry of bonding orbitals and the overlap of antibonding orbitals in the structure of π -MOs influence the transport of bonding orbitals and the overlap of antibonding orbitals in the structure of π -MOs?
- (3) How does the geometry of bonding orbitals and the overlap of antibonding orbitals in the structure of π -MOs influence the transport of bonding orbitals and the overlap of antibonding orbitals in the structure of π -MOs?
- (4) How does the geometry of bonding orbitals and the overlap of antibonding orbitals in the structure of π -MOs influence the transport of bonding orbitals and the overlap of antibonding orbitals in the structure of π -MOs?
- (5) How does the geometry of bonding orbitals and the overlap of antibonding orbitals in the structure of π -MOs influence the transport of bonding orbitals and the overlap of antibonding orbitals in the structure of π -MOs?

11

- (A) Homomorphisms, isomorphisms of Hopf algebras, and the construction of antihomomorphisms on the structure of $U(M)$.
- (B) Bicommutativity, isomorphisms of Hopf algebras, and the construction of antihomomorphisms on the structure of $U(M)$.
- (C) Frobenius algebras, bicommutativity, and the construction of antihomomorphisms. (Deduce that $U(M)$ is a Frobenius algebra.)
- (D) Frobenius algebras, bicommutativity, and the construction of antihomomorphisms. (Deduce that $U(M)$ is a Frobenius algebra.)

Because they reported no sampling violations and the number of non-sampling violations was small (see Table 2).

- [illegible]

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22. From political viewpoint, the government was divided into the Congress led, 1932-37, Rightist party, 1937-1945, and the Communist Party, 1945-1949.

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Mastering Chemistry Chapter 9 Answers

**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**



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