

### Homework 1

|   |   |
|---|---|
| <p>The average human lung expands by about 0.50 LL during each breath.</p>  | <p>Work = <math>-p \times \Delta</math><br/> <math>= -(1.0) \times (0.50\text{L})\text{atm}</math><br/> <math>= -0.50\text{L}\times\text{atm}</math><br/> Conversion of <math>1\text{L}\times\text{atm} \rightarrow \text{J} = /101.325</math><br/> <math>= -50.66</math> or rounded to <math>-51</math></p>  |
| <p>The air within a piston equipped with a cylinder absorbs 545 J of heat and expands from an initial volume of 0.11 LL to a final volume of 0.88 LL against an external pressure of 1.0 barbar.</p>  | <p><math>W = -p \times \Delta V</math>                      <math>\Delta V = 0.88 - 0.11 = .77\text{L}</math><br/> <math>= -1 \times 0.77</math><br/> <math>= -77\text{L} \times \text{atm} \times 101.325 = -78.02\text{J}</math><br/><br/> <math>\Delta E = \text{change in internal energy}</math>    <math>q = \text{heat} = 545\text{J}</math><br/> <math>w = 78.02</math><br/> <math>\Delta E = q + w</math><br/> <math>= 545 + (-78.02)</math><br/> <math>= 466.98\text{J}</math><br/> <math>= 467\text{J}</math></p>  |
| <p>A gas is compressed from an initial volume of 5.35 LL to a final volume of 1.23 LL by an external pressure of 1.00 barbar . During the compression, the gas releases 121 J of heat.</p>  | <p><math>\Delta E = q + w</math>                      <math>q = -121</math><br/> Calc work done <math>w = -p \times \Delta V</math><br/> <math>1\text{ atm} = 1.01 \times 10^5\text{ Pa}</math>, and <math>1\text{ liter} = 10^{-3}\text{ m}^3</math>, so:<br/> <math>W = (1.01 \times 10^5\text{ N/m}^2) \times (5.35 - 1.24) \times 10^{-3}\text{ m}^3</math><br/> <math>= 415.11\text{J}</math><br/><br/> So <math>\Delta E = -121 + 415.11</math><br/> <math>294</math> But with correct values its <math>291</math></p>  |
| <p>What is the work associated with decomposition of trinitrotoluene (TNT) upon detonation according to the following reaction at 281.65 KK?<br/> <math>\text{C}_7\text{H}_5\text{N}_3\text{O}_6(\text{s}) \rightarrow 3\text{N}_2(\text{g}) + 5\text{H}_2(\text{g}) + 3\text{O}_2(\text{g}) + 7\text{C}(\text{s})</math></p> | <p>1 mole <math>\text{C}_7\text{H}_5\text{N}_3\text{O}_6</math> gives <math>3/2</math> moles <math>\text{N}_2</math> and <math>5/2</math> moles <math>\text{H}_2</math> and 3 moles <math>\text{O}_2</math> gas<br/> Total gas moles produced = <math>\text{N}_2</math> moles + <math>\text{H}_2</math> moles + <math>\text{O}_2</math> moles<br/> <math>= 1.5 + 2.5 + 3 = 7</math><br/> work done = <math>-P \times dV</math> where <math>dV</math> is change in volume<br/> we use ideal gas equation <math>PV = nRT</math> ( since <math>V</math> is nothing but volume occupied by all gases)<br/> work done = <math>-nRT</math> where <math>R</math> = gas constant = <math>0.08206\text{ liter atm/molK}</math><br/> <math>= -7\text{ mol} \times 0.08206\text{ liter atm/molK} \times 281.65</math><br/> <math>= -162.8500\text{ atm liter /mol}</math><br/> we have <math>1\text{ atm liter} = 101.325\text{ J}</math><br/> hence work done = <math>-162.8500\text{ atm liter /mol} \times (101.325\text{ J/atmliter})</math><br/> <math>= -16500\text{ J/mol}</math><br/> <math>= -16.500\text{ KJ/mol}</math><br/> ( -ve sign indicates expansion work )<br/> (correct answer is <math>-16.39</math>)</p> |

# Mastering Chemistry Answers Assignment 1

**Morris Hein, Susan Arena, Cary Willard**



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