

PHET - Sugar & Salt Solutions Lab

Pre-Lab Questions (Use your notes to answer the questions below)

1. Define ionic bond: **An ionic bond is between a metal and a nonmetal.**
2. Define covalent bond: **A covalent bond is between two or more nonmetals.**
3. Define ion: **A charged particle.**

The procedure, Part A (Macro Tab)

1. Click play and download the simulation. [Sugar and Salt Solutions](#)
2. Drag the conductivity tester, labeled A in the diagram, into the beaker of water. Place the negative and positive electrodes into the water, but not touching the bottom of the beaker. When electricity is conducted by the solution, the light bulb will light up – the stronger the electric current, the brighter the bulb will glow. Record observations for the conductivity of water below.
3. Click on the salt shaker, labeled B, and drag your mouse back and forth. This will “shake” the salt into the water. As you add salt, pay attention to the light bulb on the conductivity tester. Continue to add salt until the shaker is empty. **Record observations in the chart below.**
4. Do the salt crystals “disappear” when added to the water? If they do, then they dissolve in water and are soluble. Record in the chart below. Then click the “Remove salt” button to reset the simulation.
5. Click the “sugar” bubble in the solute selector, labeled C, to switch from salt to sugar. Then repeat [step 3](#) and 4, this time using the sugar shaker. Record observations below.



Substance	Solubility in Water (yes or no)	Observations for Conductivity/Light bulb	Electrical Conductivity (yes or no)
Pure Water H_2O	N/A	Did not light up	no
Salt $NaCl$	Yes	Turned on	yes
Sugar $C_{12}H_{22}O_{11}$	Yes	Did not light up	no

Key To Lab 18

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