

Activity A:
Alpha decay

Get the Gizmo ready:

- Click **Reset** (↺).
- Check that **Alpha decay** and **Uranium** are selected.



Question: How does alpha decay change the nucleus of a radioactive atom?

1. **Predict:** As you observed in the warm-up activity, an alpha particle consists of two protons and two neutrons. How will the emission of an alpha particle affect the following?

- A. The atomic number of the atom:
B. The mass number of the atom:

decrease by two

decrease by four

2. **Calculate:** Turn on **Write equation**. What you see is an equation that shows the original uranium atom on the left. The boxes on the right represent the **daughter product**—the atom produced by radioactive decay—and the emitted alpha particle.

- A. In the top left box, write the mass number of the daughter product and press "Enter" on your keyboard. What is this number?

234

- B. In the bottom left box, write the atomic number of the daughter product and press "Enter." What is this number?

90

- C. Click on the picture below, select edit. Enter the mass number and atomic number of the alpha particle, which has the same composition as the nucleus of a helium (He) atom. After filling in the boxes in the Gizmo, write the completed equation below:



- D. According to your equation, what **isotope** remains after the alpha decay of uranium-238? (Note: You can look up element symbols on the periodic table.)

Th - 234

3. **Check:** Turn on **Show equation**, and click **Play**. The equation will appear at the end of the animation. Was your prediction correct? **Yes** If not, modify your equation above.

Key Terms Nuclear Changes Answer Key

D Keegan



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Table of Contents Key Terms Nuclear Changes Answer Key

1. Understanding the eBook Key Terms Nuclear Changes Answer Key
 - The Rise of Digital Reading Key Terms Nuclear Changes Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Key Terms Nuclear Changes Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Key Terms Nuclear Changes Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Key Terms Nuclear Changes Answer Key
 - Personalized Recommendations
 - Key Terms Nuclear Changes Answer Key User Reviews and Ratings
 - Key Terms Nuclear Changes Answer Key and Bestseller Lists
5. Accessing Key Terms Nuclear Changes Answer Key Free and Paid eBooks
 - Key Terms Nuclear Changes Answer Key Public Domain eBooks
 - Key Terms Nuclear Changes Answer Key eBook Subscription Services

- Key Terms Nuclear Changes Answer Key Budget-Friendly Options
- 6. Navigating Key Terms Nuclear Changes Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Key Terms Nuclear Changes Answer Key Compatibility with Devices
 - Key Terms Nuclear Changes Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Key Terms Nuclear Changes Answer Key
 - Highlighting and Note-Taking Key Terms Nuclear Changes Answer Key
 - Interactive Elements Key Terms Nuclear Changes Answer Key
- 8. Staying Engaged with Key Terms Nuclear Changes Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Key Terms Nuclear Changes Answer Key
- 9. Balancing eBooks and Physical Books Key Terms Nuclear Changes Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Key Terms Nuclear Changes Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Key Terms Nuclear Changes Answer Key
 - Setting Reading Goals Key Terms Nuclear Changes Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Key Terms Nuclear Changes Answer Key
 - Fact-Checking eBook Content of Key Terms Nuclear Changes Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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