

## Solutions, Acids, Bases Study Guide

### Vocabulary & Topics

Solution – solute, solvent

Polar molecule

Dissociation

Concentration – concentrated, dilute

Solubility – unsaturated, saturated, supersaturated

Alloy

Acid – Strong, weak

Base – strong, weak

Indicator, universal indicator

Hydronium ions –  $\text{H}_3\text{O}^+$

Hydroxide ions –  $\text{OH}^-$

Ph – 1-14 (7 is neutral)

Neutralization

Salts – ionic compounds formed from neutralization of acids and bases.

Practice Questions – Answers on the website

What are three factors that affect rate of dissolution?

Solute size – decreasing the size of the individual particles or grains of solute can increase surface area. Increased surface area exposes more solute to the solvent, speeding up the dissolution.

Temperature – increased temperature speeds up the molecules of the solvent, allowing more solvent to come into contact with the solute.

Stirring or mixing – mechanically mixing the solvent and solute allows more solvent to come into contact with the solute.

What is created when an acid is combined with a base?

An acid and a base can be combined in what is known as a neutralization reaction. In this reaction, the pH will become something between the pH's of the acid and the base. When they are combined, the products are water (from the  $\text{H}^+$  and  $\text{OH}^-$  ions) and an ionic compound known as a salt.

Name 3 characteristics that acids and bases share.

Acids and bases can both be

corrosive.

poisonous

electrolytes (can conduct electricity)

Both can also be described as having unique tastes and capable of changing the color of indicators.

Describe the process of dissociation of phosphoric acid ( $\text{H}_3\text{PO}_4$ ).

During the dissociation of acid, the solute ( $\text{H}_3\text{PO}_4$ ) is dissolved in water. The polar (having slightly positive and negative ends) water molecules surround the solute. The slightly negative ends of the water molecules are attracted to the positive ends of the compounds ( $\text{H}^+$ ). The slightly positive ends of the water molecules are attracted to the

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