



QUESTION PAPER 2

GRADE 9

TERM 4

ECONOMIC AND MANAGEMENT SCIENCE: ECONOMICS

TOTAL: 50 MARKS

NAME AND SURNAME: _____

CLASS: _____

QUESTION	TOPIC	MAXIMUM MARKS	LEARNER MARK	MODERATED MARK
1	VARIOUS TOPICS	10		
2	ECONOMIC SYSTEMS	10		
3	CIRCULAR FLOW	10		
4	PRICE THEORY	14		
5	SECTORS OF THE ECONOMY	6		
TOTAL		50		

Exam Agreement

I, _____ (Name & Surname), Grade _____

hereby acknowledge that I will abide by all examination rules stipulated below during the writing of my exam _____ (Subject's name).

RULES:

- I will not have any electronic equipment in my possession (cell phones, smart watches and earphones).
- I will not cause any disruption during this exam in the exam venue.
- I will be honest at all times.

All rules, as discussed with me, will be honoured and obeyed.

Learner Signature _____

QUESTION 1

Various options are provided as possible answers for the following questions. Choose the correct answer and write the letter (A-C) next to the question number (1.1.1 – 1.1.5).

1.1.1 Which of the following statements is true about a mixed economy?

- The government takes all the decisions.
- The economy is a combination of state and private sector involvement.
- Private hospitals do not exist.
- The government is not involved in the economy at all.

1.1.2 DUKKIE Mines will operate in the ... sector of the economy.

- primary
- secondary
- tertiary
- mining

1.1.3 If the price of a good rises and the quantity demanded will ...

- increase
- decrease
- stay the same
- increase the supply of the product

1.1.4 Which of the following is an example of a factor of production?

- A cell phone
- A factory building
- A new computer
- A loaf of bread

1.1.5 The primary purpose of taxes in the circular flow of the economy is to ...

- generate income for the government
- control consumer spending
- reduce inflation
- pay government employees

(5)

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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the ion meta dissolves

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conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away cooper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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iron with copper ii sulfate i obtained the results to this experiment through adding heated CuSO_4 solution to the iron powder stirring the solution adding water then removing water adding acetone heating the acetone on low

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