



User Manual

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Elix 5 Manual

Léo Pomar



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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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g al x 1 mol al x 2 mol al₂O₃ x 101 964 g al₂O₃ 302 3 g al₂O₃ 26 982 g al 4 mol al 1 mol al₂O₃ c determine the percent yield
 yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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 weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12
 c is the limiting reagent 4 33 g of h 2 are left over 13

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 yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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 ch₃oh can be produced by the following reaction 2h₂ co ch₃oh calculate the theoretical yield of ch₃oh if 68 5 g of co is
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 what is the

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