

(Question 1 continued)

- (d) The cross-sectional area of the tube is $1.3 \times 10^{-3} \text{ m}^2$ and the temperature of air is 300 K. Estimate the number of moles of air in the tube. [2]

$$\begin{aligned}
 PV &= nRT & (b) A &= PY & \text{or } n &= \frac{LA}{RT} \\
 n &= \frac{PV}{RT} & & & & n = \frac{(5 \times 10^4)(1.3 \times 10^{-3})}{(8.31)(300)} \\
 \text{but you are given Area} &= 1.3 \times 10^{-3} \text{ m}^2 & & & & n = 0.026 \text{ mol} \\
 \text{since } \rho &= \frac{P}{V_H} \text{ or } P = \rho H & & & & \approx 0.03 \text{ mol}
 \end{aligned}$$

- (e) The equation in (b) may be used to predict the pressure of the air at extremely large values of $\frac{1}{H}$. Suggest why this will be an unreliable estimate of the pressure. [2]

Very large $1/H$ means small volume. Small volume means high pressure.

In ideal gas law, it stated that molecules are not close to each other and hence the forces between them are negligible.

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