

## STUDY GUIDE: Understanding Properties of Matter

1. Use examples to describe the difference between physical and chemical properties of matter. (See handout)
2. List 9 physical properties of matter with an example of each. **Density, Ductility, Electrical Conductivity, Magnetism, Malleability, Solubility, Specific Heat, State, Thermal Conductivity** (Examples will vary but should be similar to the examples found in the Properties of Matter Booklet.)
3. List 2 chemical properties of matter with an example of each. **Flammability, Reactivity** (Examples will vary but should be similar to the examples found in the Properties of Matter Booklet.)
4. Which property of matter is observed when water changes from a solid to a liquid? **State**
5. Adding food coloring to the solution is an example of **solubility**.
6. If bubbles start to rise from a solution in the scientist's test tube shown, it would be an indication of **reactivity**.
7. What "point" is observed when water changes from a solid to a liquid? **Melting point**
8. Which property(ies) would you use to describe most metals? **Ductility or Malleability or Electrical Conductivity**
9. Which of the properties of metals means that it can be drawn into wires? **Ductility**
10. Which property allows metals to be bent rather than broken? **Malleability**
11. Besides being valuable, gold is special because it can be hammered into a very thin foil. This property is called **malleability**.

Use the table below to interpret questions 12 – 15.

| SUBSTANCE        | DENSITY (g/cm <sup>3</sup> ) | SUBSTANCE      | DENSITY (g/cm <sup>3</sup> ) |
|------------------|------------------------------|----------------|------------------------------|
| Aluminum (solid) | 2.7                          | Ice (solid)    | 0.93                         |
| Pyrite (solid)   | 5.02                         | Water (liquid) | 1.00                         |
| Mercury (liquid) | 13.55                        | Zinc (solid)   | 7.13                         |
| Silver (solid)   | 10.50                        | Wood (oak)     | 0.85                         |

12. What substance has a density more than 13 times greater than water? **Mercury**
13. What is the density of ice? **.93g/cm<sup>3</sup>**
14. Which metal is a liquid at room temperature? **Mercury**
15. What two variables determine the density of a substance? **Mass/volume**
16. The electromagnets found inside of doorbells are an example of which properties of matter? **Magnetism; Electrical conductivity; Ductility (coiled wire around metal)**
17. Chemical weathering, such as acid rain causing a statue to fizz and deteriorate is an example of which property of matter? **Reactivity**
18. Which property of matter is exemplified by pressing metals into coins? **Malleability**
19. Which property of matter is observed when water changes from a liquid to a solid? **State**
20. What "point" is observed when water changes from a liquid to a solid? **Freezing point**
21. Which property of matter is observed when water changes from a liquid to a gas? **State**
22. What "point" is observed when water changes from a liquid to a gas? **Boiling point**

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