Reviewing Content

- The solvent is the substance in which the solute is dissolved.
- Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
- 44. solubility: the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.

saturated solution: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure, unsaturated solution: a solution that contains less solute than a saturated solution at a given temperature and pressure, miscible: describes liquids that dissolve in each other, immiscible: describes liquids that are insoluble in each other.

- 45. Particles of solute crystalline.
- No; if there were undissolved solute, the excess solute would come out of a supersaturated solution.
- 47. 5.55 × 10² g AgNO₃
- 48. Solubility increases with pressure.
- 49. a. 1.6 × 10⁻² g/L b. 4.7 × 10⁻² g/L
- Dilute and concentrated are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
- Molarity is the number of moles of solute dissolved in one liter of solution.
 - a. 1.3M KCI
 - b. 3.3 × 10⁻¹M MgCl₂
- 52. 2.00 × 10³ mL
- a. 5.0 ×10⁻¹ mol NaCl, 29 g NaCl
 b. 1.0 mol KNO₃, 1.0 × 10⁻² g KNO₃
 c. 2.5 × 10² mol CaCl₂, 2.8 g CaCl₂
- a. 2.3 × 10¹ g NaCl
 b. 2.0 g MgCl₂
- a. 16% (v/v) ethanol
 b. 63.6% (v/v) isopropyl alcohol

- 56. Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation, freezing-point depression, and vaporpressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution; relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
- 57. a. sea water
 - b. 1.554 KNO₂
 - c. 0.100M MgCl₂
- The effective molality of the Ca(NO₃)₂ solution is 3m. The effective molality of the NaNO₃ solution is 2m.
- 59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
- The salt lowers the freezing point of the ice-water cooling mixture.
- 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
- 62. Add 27.0 g H₂O to 32.0 g CH₂OH.
- 63. m. 100.25°C
 - b. 101.54°C
- 64. m. -1.46°C
 - b. -2.2°C
- 65. a. -1.1°C
 - ba. -0.747C
 - c. -1.5°C

Understanding Concepts

66. a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

Mastering Chemistry Answers Chapter 5

Kurt W. Kolasinski

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Table of Contents Mastering Chemistry Answers Chapter 5

- 1. Understanding the eBook Mastering Chemistry Answers Chapter 5
 - The Rise of Digital Reading Mastering Chemistry Answers Chapter 5
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Mastering Chemistry Answers Chapter 5
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mastering Chemistry Answers Chapter 5
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Mastering Chemistry Answers Chapter 5
 - Personalized Recommendations
 - Mastering Chemistry Answers Chapter 5 User Reviews and Ratings
 - Mastering Chemistry Answers Chapter 5 and Bestseller Lists

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 - Mastering Chemistry Answers Chapter 5 Enhanced eBook Features
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 - Highlighting and Note-Taking Mastering Chemistry Answers Chapter 5
 - Interactive Elements Mastering Chemistry Answers Chapter 5
- 8. Staying Engaged with Mastering Chemistry Answers Chapter 5
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - $\circ\,$ Following Authors and Publishers Mastering Chemistry Answers Chapter 5
- 9. Balancing eBooks and Physical Books Mastering Chemistry Answers Chapter 5
 - Benefits of a Digital Library
 - \circ Creating a Diverse Reading Collection Mastering Chemistry Answers Chapter 5
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mastering Chemistry Answers Chapter 5
 - Setting Reading Goals Mastering Chemistry Answers Chapter 5
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mastering Chemistry Answers Chapter 5
 - Fact-Checking eBook Content of Mastering Chemistry Answers Chapter 5
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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