

Reviewing Content

42. The solvent is the substance in which the solute is dissolved.
43. Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
44. **solubility:** the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.
saturated solution: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. **unsaturated solution:** a solution that contains less solute than a saturated solution at a given temperature and pressure. **miscible:** describes liquids that dissolve in each other. **immiscible:** describes liquids that are insoluble in each other.
45. Particles of solute crystallize.
46. No; if there were undissolved solute, the excess solute would come out of a supersaturated solution.
47. 5.55×10^2 g AgNO_3
48. Solubility increases with pressure.
49. a. 1.6×10^{-2} g/L
b. 4.7×10^{-2} g/L
50. *Dilute* and *concentrated* are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
51. Molarity is the number of moles of solute dissolved in one liter of solution.
a. 1.3M KCl
b. 3.3×10^{-1} M MgCl_2
52. 2.00×10^1 mL
53. a. 5.0×10^{-1} mol NaCl, 29 g NaCl
b. 1.0 mol KNO_3 , 1.0×10^{-2} g KNO_3
c. 2.5×10^2 mol CaCl_2 , 2.8 g CaCl_2
54. a. 2.3×10^1 g NaCl
b. 2.0 g MgCl_2
55. a. 16% (v/v) ethanol
b. 63.6% (v/v) isopropyl alcohol
56. Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation, freezing-point depression, and vapor-pressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution; relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
57. a. sea water
b. 1.5M KNO_3
c. 0.100M MgCl_2
58. The effective molality of the $\text{Ca}(\text{NO}_3)_2$ solution is 3m. The effective molality of the NaNO_3 solution is 2m.
59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
60. The salt lowers the freezing point of the ice-water cooling mixture.
61. 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
62. Add 27.0 g H_2O to 32.0 g CH_3OH .
63. a. 100.26°C
b. 101.54°C
64. a. -4.46°C
b. -2.2°C
65. a. -1.1°C
b. -0.74°C
c. -1.5°C

Understanding Concepts

66. a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

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R Barnett

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Table of Contents Mastering Chemistry Answers Chapter 17

1. Understanding the eBook Mastering Chemistry Answers Chapter 17
 - The Rise of Digital Reading Mastering Chemistry Answers Chapter 17
 - Advantages of eBooks Over Traditional Books
2. Identifying Mastering Chemistry Answers Chapter 17
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mastering Chemistry Answers Chapter 17
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mastering Chemistry Answers Chapter 17
 - Personalized Recommendations

- Mastering Chemistry Answers Chapter 17 User Reviews and Ratings
- Mastering Chemistry Answers Chapter 17 and Bestseller Lists
- 5. Accessing Mastering Chemistry Answers Chapter 17 Free and Paid eBooks
 - Mastering Chemistry Answers Chapter 17 Public Domain eBooks
 - Mastering Chemistry Answers Chapter 17 eBook Subscription Services
 - Mastering Chemistry Answers Chapter 17 Budget-Friendly Options
- 6. Navigating Mastering Chemistry Answers Chapter 17 eBook Formats
 - ePub, PDF, MOBI, and More
 - Mastering Chemistry Answers Chapter 17 Compatibility with Devices
 - Mastering Chemistry Answers Chapter 17 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mastering Chemistry Answers Chapter 17
 - Highlighting and Note-Taking Mastering Chemistry Answers Chapter 17
 - Interactive Elements Mastering Chemistry Answers Chapter 17
- 8. Staying Engaged with Mastering Chemistry Answers Chapter 17
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mastering Chemistry Answers Chapter 17
- 9. Balancing eBooks and Physical Books Mastering Chemistry Answers Chapter 17
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mastering Chemistry Answers Chapter 17
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mastering Chemistry Answers Chapter 17
 - Setting Reading Goals Mastering Chemistry Answers Chapter 17
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mastering Chemistry Answers Chapter 17
 - Fact-Checking eBook Content of Mastering Chemistry Answers Chapter 17

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- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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