

Model 2: A Proposed Two-Step Mechanism for a Chemical Reaction.



Experimental rate law: $\text{rate} = k_{\text{exp}} (\text{NO})^2 (\text{O}_2)$

Proposed Mechanism:

Step	Molecular Event	Rate of Forward Step	Relative Rate
1	$\text{NO(g)} + \text{NO(g)} \rightleftharpoons \text{N}_2\text{O}_2\text{(g)}$	$k_1 (\text{NO})^2$	fast equilibrium
2	$\text{N}_2\text{O}_2\text{(g)} + \text{O}_2\text{(g)} \rightleftharpoons 2 \text{NO}_2\text{(g)}$	$k_2 (\text{N}_2\text{O}_2) (\text{O}_2)$	slow forward

Rate law for the proposed mechanism = rate for step 2 = $k_2 (\text{N}_2\text{O}_2) (\text{O}_2)$

Critical Thinking Questions

For CTQs 8 - 11, answer the questions individually and then discuss with your teammates to reach consensus.

8. For this mechanism, which forward steps are unimolecular and which forward steps are bimolecular?

$\text{Unimolecular} = 1, 2$
 $\text{Bimolecular} = 1$

9. Confirm that the "rate of forward step" given for each step is consistent with your answer to CTQ 3.

10. Why is the rate law for the proposed mechanism approximately equal to the rate of forward step 2?

11. Based on the analysis thus far, does this proposed mechanism appear to be consistent with the experimental rate law for the overall reaction in Model 1. Explain your reasoning.

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