

The Mole and Equations

This is the moment where the 'mole of moles' comes in. It's the equation from page 104 that makes it all work.

You Can Use Moles to Calculate Masses in Reactions

Remember those balanced equations back on page 107? Well, the **balancing** numbers (the chemical formulae of the reactants and products) tell you how much of each substance takes part in or is formed during the reaction.

For example:



In this reaction, **1 mole** of magnesium and **2 moles** of hydrochloric acid react together to form **1 mole** of magnesium chloride and **1 mole** of hydrogen gas.

You Can Balance Equations Using Reacting Masses

If you know the **masses** of the **reactants** and **products** that took part in a reaction, you can work out the **balanced symbol equation** for the reaction. Here are the steps you should take:

1. Write the **mass** of each substance by its **relative formula mass** for that the **number of moles**.
2. Write the **number of moles** of each substance by the **symbol number of moles** in the reaction.
3. If any of the numbers aren't whole numbers, multiply all the numbers by the same amount so that they all become whole numbers.
4. Write the **balanced symbol equation** for the reaction by putting these numbers in front of the chemical formulae.

EXAMPLE 1

81 g of zinc with CuSO_4 reacts completely with 0.80 g of carbon to form 32 g of carbon dioxide and 65 g of zinc. Write a balanced symbol equation for this reaction.



1. Write out the **mass** of each substance in the reaction.

2. Write the **number of moles** of each substance by its **symbol number of moles**.

3. If any of the numbers aren't whole numbers, multiply all the numbers by the same amount so that they all become whole numbers.

4. Write the **balanced symbol equation** for the reaction by putting these numbers in front of the chemical formulae.



Where do moles live? Digaburrow...

The calculations on this page have lots of steps, so the best way to learn how to do them is to do them in the practical. Luckily, there are some pointers below to get you started. Don't say I didn't warn you. Start off working on the first two.

1. 81 g of Zn reacts completely with 16 g of CuSO_4 to produce 152 g of ZnSO₄.
2. 12 g of C reacts completely with 32 g of O_2 to produce 44 g of CO_2 .

3. Calculate how many moles of each substance involved in each reaction.
4. Use your answers to put up a balanced symbol equation for the reaction.

(1 mark)
(2 marks)

Limiting Reactants

Reactions don't go on forever — you need stuff in the reaction that can react. If one reactant gets **used up** in a reaction before the rest, then the reaction will **stop**. That reactant's called **limiting**.

Reactions Stop When One Reactant Is Used Up

When some **magnesium carbonate** (MgCO_3) is placed into a beaker of **hydrochloric acid**, you can tell a **reaction** is taking place because you see lots of **bubbles of gas** being given off. After a while, the amount of **acid** **runs down** and the reaction **eventually stops**.

1. The reaction stops when all of one of the reactants is **used up**. Any other reactants are in **excess**. They're usually added in excess to **make sure** that the other reactant is used up.
2. The reactant that's **used up** in a reaction is called the **limiting reactant** (because it limits the amount of product that's formed).
3. The amount of product formed is **directly proportional** to the amount of **limiting reactant**. For example, if you **halve** the amount of limiting reactant the amount of product formed will also **halve**. If you **double** the amount of limiting reactant the amount of product will **double** (as long as it is still the limiting reactant).
4. This is because if you add **more reactant** there will be **more product particles** to take part in the reaction, which means **more product particles**.



The Amount of Product Depends on the Limiting Reactant

You can calculate the **mass of a product** formed in a reaction by using the **mass of the limiting reactant** and the **balanced reaction equation**.

1. Write out the **balanced equation**.
2. Work out **relative formula masses** (M_r) of the reactant and product you want.
3. Find out **how many moles** there are of the substance you **know** the mass of.
4. Use the balanced equation to work out **how many moles** there'll be of the **other** substance. In this case, that's how many moles of **product** will be made of this many moles of **reactant**.
5. Use the number of moles to calculate the **mass**.

EXAMPLE 2

Calculate the mass of aluminium oxide formed when 105 g of aluminium is burned in air.

1. Write out the **balanced equation**: $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
2. Calculate the **relative formula masses**: $4 \times 27 = 108$, $3 \times 32 = 96$, $2 \times 102 = 204$
3. Calculate the **number of moles** of aluminium = 105 g
4. Look at the **ratio of moles** in the equation: $4 \text{ moles of Al} \rightarrow 2 \text{ moles of Al}_2\text{O}_3$
5. Calculate the **mass** of 2.5 moles of aluminium oxide

'A Dash of Mad to the Head' — my favourite album by Moby-Dick

I've read it before, I'll read it again! — practical moles problem. So before you get down to work on the questions below, make an album off to have a cup of tea, have a go at the questions below.

1. The balanced equation for the reaction between chlorine and potassium bromide is: $\text{Cl}_2 + 2\text{KBr} \rightarrow 2\text{KCl} + \text{Br}_2$
2. Calculate the mass of potassium chloride produced when 23.8 g of potassium bromide reacts with excess chlorine. $A_r(\text{K}) = 39$, $A_r(\text{Br}) = 80$, $A_r(\text{Cl}) = 35.5$

(1 mark)
(2 marks)

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Paddy Gannon



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