

Greatest Common Factor (A)

Name: _____

Date: _____

Use the prime factors of the numbers in each set to calculate the greatest common factor.

a) $64 = \textcircled{2} \times \textcircled{2} \times \textcircled{2} \times 2 \times 2 \times 2$

$72 = \textcircled{2} \times \textcircled{2} \times \textcircled{2} \times 3 \times 3$

$\text{GCF} = \textcircled{2} \times \textcircled{2} \times \textcircled{2} = 8$

b) 96

48

c) 20

8

d) 36

90

e) 60

44

f) 12

56

g) 28

8

h) 90

9

i) 72

96

j) 72

66

Greatest Common Factor Kuta

R Bogdan

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Greatest Common Factor Kuta:

Eventually, you will no question discover a further experience and ability by spending more cash. nevertheless when? attain you assume that you require to get those every needs in the manner of having significantly cash? Why dont you attempt to acquire something basic in the beginning? Thats something that will guide you to understand even more on the subject of the globe, experience, some places, taking into consideration history, amusement, and a lot more?

It is your completely own era to take steps reviewing habit. along with guides you could enjoy now is **Greatest Common Factor Kuta** below.

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web donor whereas a base generically represented by b is a proton acceptor when acids and bases combine they swap a proton $ha + b \rightleftharpoons h^+ + b^-$ strong acids e.g. sulfuric acid have a very pronounced tendency to transfer a proton to a base whereas weak acids

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web understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web 1 approximately 3 g of hydrated oxalic acid $H_2C_2O_4 \cdot 2H_2O$ was weighted to the nearest 0.01 g in a 50 mL beaker 2 about

30 ml of distilled water was added to the beaker to dissolve the oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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