Determining the Mole Ratios in a Chemical Reaction

A fealinged disensed reaction equation gives the mole ratios of the reactions and the products as coefficients. When some of the chemical formulas are not known, an experiment must be conducted to help determine the mole ratios.

This experiment uses two substances, an unidentified acid and codium hydroxide Q4xQ4Q, which react according to the unbalanced and incomplete reaction equation below:

It is possible to identify the coefficients, a and b, for the reactants, without knowing the identity of the acid and the products of the reaction. The process that you will use to determine the coefficients is called coefficients without variations. You will prepare a series of mightires of the two reactants. Each mixture will have the name total volume and the name total number of moles of reactants. The reaction is excitement, thus the mixture that generates the most heat energy will be the reaction that completely consumes both the unidentified acid and the podum hydroxide. By analyzing the results of the series of reactants.

Thus will also be determining whether the unatentified and in monoprotic, diprotic, or improtice. Moreoprotic analyses able to denote one proton per molecule. Common examples of monoprotice ands are hydrochiloric and (BCD) and acetic and ($BC_2H_2O_2$). Polyprotic acetic, in contrast, can denote more than one proton per molecule. Sulfamo and $O(pCO_2)$ and carbonic acid $O(pCO_3)$ are common algorithm acids that are able to donate two protons per molecule. Phosphonic acid $O(pCO_3)$ are common deposits acid $O(pCO_3)$ are common provide acids.

OBJECTIVES

- Measure the temperature changes of a series of reactions.
- Determine the stoodsometry of an oxidation-reduction reaction is which the reactions are known but the products are unknown.
- Determine whether the unidentified acid is monoprotic, digestic, or triprotic.

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G Orfield

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