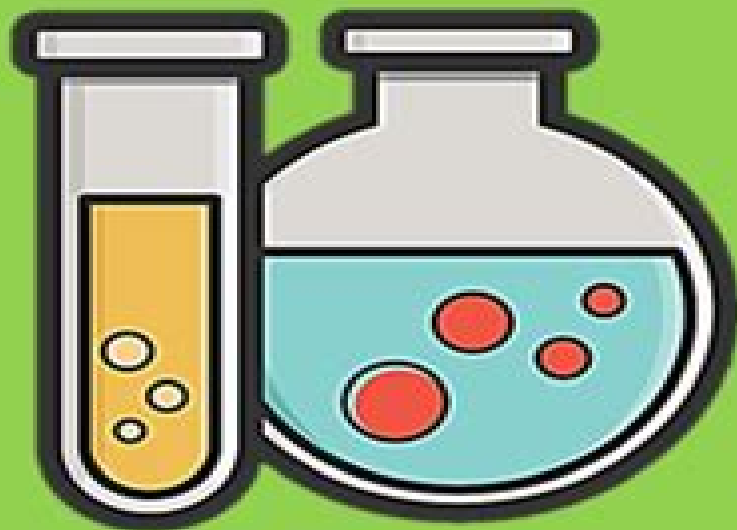


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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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1 K^+ PO_4^{3-} aq Al^{3+} aq PO_4^{3-} s 3K^+ PO_4^{3-} aq 3Cl^- aq Al^{3+} aq 3NO_3^- aq AlPO_4 s 3K^+ aq PO_4^{3-} aq AlPO_4 s 2Fe^{3+} aq 3Mg s

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver i ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13 $2\text{ag aq} + 2\text{f aq} + 2\text{nh}_4 \text{ aq} + \text{cr}_2\text{o}_7^{2-} \text{ aq} \rightarrow \text{ag}_2\text{cr}_2\text{o}_7 \text{ s} + 2\text{nh}_4 \text{ aq} + 2\text{f aq}$ because two nh 4 aq

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