

Classifying and naming compounds

Classify each compound as ionic or covalent and then give the appropriate name for the compound.

Compound Formula	Ionic or Covalent	Compound Name	Molecular Mass
MgCl ₂	ionic	Magnesium Chloride	95.21
SO ₃ ²⁻	ionic	sulfite ion	80.06
P ₂ S ₃	covalent	Diphosphorus Trisulfide	158.1
Al ₂ (SO ₃) ₃	ionic	Aluminum sulfite	234.15
P ₂ O ₅	covalent	Diphosphorus Pentoxide	141.94
KNO ₃	ionic	Potassium Nitrate	101.1
Na ₂ CrO ₄	ionic	Sodium Chromate	161.97
SF ₄	covalent	Sulfur Tetrafluoride	108
N ₂ O ₄	covalent	Dinitrogen Tetraoxide	92.01
CO	covalent	Carbon monoxide	28
MgCO ₃	ionic	Magnesium Carbonate	84.3
SO ₂	covalent	Sulfur Dioxide	64
PCl ₅	covalent	Phosphorus Pentachloride	208.24
Al ₂ O ₃	ionic	Aluminum Oxide	101.96
P ₂ Br ₃	covalent	Phosphorus Tribromide	270.68
KMnO ₄	ionic	Potassium Permanganate	158
SiF ₄	covalent	Silicon Tetrafluoride	104.07
SF ₂	covalent	Sulfur Difluoride	70.06
N ₂ O ₄	covalent	Dinitrogen Tetraoxide	92
CO ₃ ²⁻	ionic	Carbonate ion	60
MgI ₂	ionic	Magnesium Iodide	278.11
SO ₄ ²⁻	ionic	sulfate ion	96
PI ₃	covalent	Phosphorus Triiodide	665.49
Al ₂ (SO ₄) ₃	ionic	Aluminum Sulfate	150
PO ₃	covalent	Phosphorus Trioxide	78.97
KBr	ionic	Potassium Bromide	119
NaI	ionic	Sodium Iodide	149.87
SF ₆	covalent	Sulfur hexafluoride	146
N ₄ O ₁₀	covalent	Tetra Nitrogen Decaoxide	216
CO ₂	covalent	Carbon Dioxide	44

Al₂S₃

Ionic Compounds Content Mastery Answers

Daniel F McAuley



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