**Hess's Law Practice Problems** 

\*\*\*Please check back for the answers. I will repost this document with the answers at a later date\*\*\*

1. Given the following equations and  $\Delta H^*$  values, determine the heat of reaction (kJ) at 298 K for the reaction:

$$B_2H_4(g) + 6 Cl_2(g) \rightarrow 2 BCl_3(g) + 6 HCl(g)$$

 $BC1_1(g) + 3 H_2O(1) \longrightarrow H_2BO_1(g) + 3 HC1(g)$   $B_1H_1(g) + 6 H_2O(1) \longrightarrow 2 H_2BO_1(g) + 6 H_2(g)$  $1/2 H_2(g) + 1/2 C1_2(g) \longrightarrow BC1(g)$   $\Delta H^{o}/kJ = -112.5$   $\Delta H^{o}/kJ = -493.4$  $\Delta H^{o}/kJ = -92.3$ 

Answer: -1376 kJ

2. Given the following equations and  $\Delta H^*$  values, determine the heat of reaction (kJ) at 298 K for the reaction:

$$2 \text{ OF}_2(g) + 2 \text{ S(s)} \rightarrow \text{SO}_2(g) + \text{SF}_4(g)$$

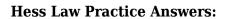
 $OF_2(g) + H_2O(1) \longrightarrow O_2(g) + 2 HF(g)$   $SF_4(g) + 2 H_2O(1) \longrightarrow 4 HF(g) + SO_2(g)$  $S(s) + O_2(g) \longrightarrow SO_2(g)$   $\Delta H^{+}/kJ = -276.6$   $\Delta H^{+}/kJ = -827.5$  $\Delta H^{+}/kJ = -296.9$ 

Answer: -319.5 kJ

3. Determine  $\triangle H^*/kJ$  for the following reaction using the listed enthalpies of reaction:

# **Hess Law Practice Answers**

Barron's Educational Series, Mark Kernion, Joseph A. Mascetta



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