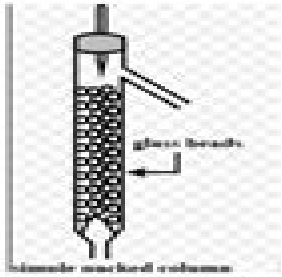
Suggest why chloroethanoic acid is stronger than ethanoic acid	(this question is already discussed a little further down.)		
fill in rate table	1.00 0.16 1.35 I beg someone write what the values were if they remember	K=250 Yes i agree - needs to be 1.4 as all values were to two sig figs Weren't we they to 3 dp though? The rate of the first one was 1.00 so I'm guessing 3sf of 2dp	3
empirical formula	C6H10O3S2	Percentages dudbt add up to 100 fuck me Didn't even see the sulphur, fml same same i thought i meant i was in the fuel area. YAY i got this!!!	3
the labelling distillation	for the tube it was fractionating column or glass beads the bottom of the condenser was water in and top was water out	 reject anti bumping granules might accept silica BEADS	3
Draw a line for the reaction with chloroethanonic acid Suggest why chloroethanoic acid is a stronger acid than ethanoic acid.	Was the line of chloroethanoic acid to the left of the first line, but levelled off at the same point? Yes but it levelled out before, same height though. Ah ok. Dissociates more Chlorine is more electronegative than hydrogen so will withdraw electron density form the carboxyl group meaning that the O H bond breaks more easily dissociating into H+ making the solution more acidic	Cl is electron withdrawing, it polarises the bond, weakens the O-H bond which releases more H+ than ethanoic acid Neg inductive effect? Chloroethanoic acid forms a more stable conjugate base?	3
NMR Question?	Use the integration values and the ratios to identify the structure at two peaks. Draw this structure.	Was there a typo though? Regarding what? No dont think so, it was just because 2 of the peaks were the same environment almost so people got thrown off	5 2 1 2

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Daniel F McAuley

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Table of Contents F335 June 2013 Unofficial Mark Scheme

1. Understanding the eBook F335 June 2013 Unofficial Mark Scheme
    - The Rise of Digital Reading F335 June 2013 Unofficial Mark Scheme
    - Advantages of eBooks Over Traditional Books
2. Identifying F335 June 2013 Unofficial Mark Scheme
    - Exploring Different Genres
    - Considering Fiction vs. Non-Fiction
    - Determining Your Reading Goals
3. Choosing the Right eBook Platform
    - Popular eBook Platforms
    - Features to Look for in an F335 June 2013 Unofficial Mark Scheme
    - User-Friendly Interface
4. Exploring eBook Recommendations from F335 June 2013 Unofficial Mark Scheme
    - Personalized Recommendations
    - F335 June 2013 Unofficial Mark Scheme User Reviews and Ratings
    - F335 June 2013 Unofficial Mark Scheme and Bestseller Lists

5. Accessing F335 June 2013 Unofficial Mark Scheme Free and Paid eBooks
    - F335 June 2013 Unofficial Mark Scheme Public Domain eBooks
    - F335 June 2013 Unofficial Mark Scheme eBook Subscription Services
    - F335 June 2013 Unofficial Mark Scheme Budget-Friendly Options
6. Navigating F335 June 2013 Unofficial Mark Scheme eBook Formats
    - ePub, PDF, MOBI, and More
    - F335 June 2013 Unofficial Mark Scheme Compatibility with Devices
    - F335 June 2013 Unofficial Mark Scheme Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of F335 June 2013 Unofficial Mark Scheme
    - Highlighting and Note-Taking F335 June 2013 Unofficial Mark Scheme
    - Interactive Elements F335 June 2013 Unofficial Mark Scheme
8. Staying Engaged with F335 June 2013 Unofficial Mark Scheme
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers F335 June 2013 Unofficial Mark Scheme
9. Balancing eBooks and Physical Books F335 June 2013 Unofficial Mark Scheme
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection F335 June 2013 Unofficial Mark Scheme
10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
11. Cultivating a Reading Routine F335 June 2013 Unofficial Mark Scheme
    - Setting Reading Goals F335 June 2013 Unofficial Mark Scheme
    - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of F335 June 2013 Unofficial Mark Scheme
    - Fact-Checking eBook Content of F335 June 2013 Unofficial Mark Scheme
    - Distinguishing Credible Sources
13. Promoting Lifelong Learning

    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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