



26. Gases W and X react in a closed, rigid vessel to form gases Y and Z according to the equation above. The initial pressure of $W_{(g)}$ is 1.20 atm and that of $X_{(g)}$ is 1.60 atm. No $Y_{(g)}$ or $Z_{(g)}$ is initially present. The experiment is carried out at constant temperature. What is the partial pressure of $Z_{(g)}$ when the partial pressure of $W_{(g)}$ has decreased to 1.0 atm?

- (A) 0.20 atm (B) 0.40 atm (C) 1.0 atm
(D) 1.2 atm (E) 1.4 atm

$$\begin{array}{r}
 W + X \rightarrow Y + Z \\
 1.2 \quad 1.6 \quad 0 \quad 0 \\
 - 1.2 \\
 \hline
 1.0
 \end{array}$$

27. On a mountaintop, it is observed that water boils at 90°C not at 100°C as at sea level. This phenomenon occurs because on the mountaintop the

- (A) equilibrium water vapor pressure is higher due to the higher atmospheric pressure
(B) equilibrium water vapor pressure is lower due to the higher atmospheric pressure
(C) equilibrium water vapor pressure equals the atmospheric pressure at a lower temperature
(D) water molecules have a higher average kinetic energy due to the lower atmospheric pressure
(E) water contains a greater concentration of dissolved gases

28. A 40.0 mL sample of 0.25 M KOH is added to 60.0 mL of 0.15 M $\text{Ba}(\text{OH})_2$. What is the molar concentration of $\text{OH}^-_{(aq)}$ in the resulting solution? (Assume that the volumes are additive.)

- (A) 0.10 M (B) 0.19 M (C) 0.28 M
(D) 0.40 M (E) 0.55 M

$$\begin{array}{l}
 .04 \times .25 = .01 \\
 .06 \times .15 = .009 \\
 \hline
 .019 / .1 = .19
 \end{array}$$



29. A 0.03 mol sample of $\text{NH}_4\text{NO}_3(s)$ is placed in a 1 L evacuated flask, which is then sealed and heated. The $\text{NH}_4\text{NO}_3(s)$ decomposes completely according to the balanced equation above. The total pressure in the flask measured at 400 K is closest to which of the following? (The value of the gas constant, R , is $0.082 \text{ L atm mol}^{-1} \text{ K}^{-1}$.)

- (A) 3 atm (B) 1 atm (C) 0.5 atm
(D) 0.1 atm (E) 0.03 atm

$$(1)(P) = .09(.082)(400)$$



30. For the reaction of ethylene represented above, ΔH is $-1,323 \text{ kJ}$. What is the value of ΔH if the combustion produced liquid water $\text{H}_2\text{O}(l)$, rather than water vapor $\text{H}_2\text{O}(g)$? (ΔH for the phase change $\text{H}_2\text{O}(g) \rightarrow \text{H}_2\text{O}(l)$ is -44 kJ mol^{-1} .)

- (A) $-1,235 \text{ kJ}$ (B) $-1,279 \text{ kJ}$ (C) $-1,323 \text{ kJ}$
(D) $-1,367 \text{ kJ}$ (E) $-1,411 \text{ kJ}$

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