

Lehninger Biochemistry Chapter 2 Exam With Verified Questions With Answers

Which of these statements about hydrogen bonds is not true?

- A) Hydrogen bonds account for the anomalously high boiling point of water.
- B) In liquid water, the average water molecule forms hydrogen bonds with three to four other water molecules.
- C) Individual hydrogen bonds are much weaker than covalent bonds.
- D) Individual hydrogen bonds in liquid water exist for many seconds and sometimes for minutes.
- E) The strength of a hydrogen bond depends on the linearity of the three atoms involved in the bond. - Answer D) Individual hydrogen bonds in liquid water exist for many seconds and sometimes for minutes.

A true statement about hydrophobic interactions is that they:

- A) are the driving force in the formation of micelles of amphipathic compounds in water.
- B) do not contribute to the structure of water-soluble proteins.
- C) have bonding energies of approximately 20-40 Kjoule per mole.
- D) involve the ability of water to denature proteins.
- E) primarily involve the effect of polar solutes on the entropy of aqueous systems. - Answer A) are the driving force in the formation of micelles of amphipathic compounds in water.

Hydrophobic interactions make important energetic contributions to:

- A) binding of a hormone to its receptor protein.
- B) enzyme-substrate interactions.
- C) membrane structure.
- D) three-dimensional folding of a polypeptide chain.
- E) all of the above are true. - Answer E) all of the above are true.

Dissolved solutes alter some physical (colligative) properties of the solvent water because they change the:

- A) concentration of the water.
- B) hydrogen bonding of the water.
- C) ionic bonding of the water.
- D) pH of the water.
- E) temperature of the water. - Answer A) concentration of the water.

Osmosis is movement of a:

- A) charged solute molecule (ion) across a membrane.

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