

## Worksheet – Gas Laws III

### Gas Laws: Combined Gas Law

The following problems may be solved using the relationships described by either Charles' or Boyles' laws, together in the combined gas law. Show all the steps to the solution of the problem.

$$V_1P_1/T_1 = V_2P_2/T_2$$

1. Carbon dioxide occupies a 2.54 dm<sup>3</sup> container at STP. What will be the volume when the pressure is 150 KPa and 26°C?

$$V_2 = \frac{[2.54 \text{ dm}^3][101.3 \text{ KPa}][26^\circ\text{C} + 273\text{K}]}{[150 \text{ KPa}][273 \text{ K}]} = 1.88 \text{ dm}^3$$

2. Oxygen occupies a fixed container of 5.5L at STP. What will happen to the pressure if the temperature rises to 300K?

$$P_2 = \frac{[5.5 \text{ L}][101.3 \text{ KPa}][300 \text{ K}]}{[5.5 \text{ L}][273 \text{ K}]} = 111.3 \text{ KPa}$$

3. Methane is compressed in a closed 15.8 dm<sup>3</sup> container at 101.3 KPa. If the volume drops to 8.7 dm<sup>3</sup> and the temperature begins at 25°C and then drops to 18°C , what will the pressure of the gas be?

$$P_2 = \frac{[101.3 \text{ KPa}][15.8 \text{ dm}^3][300 \text{ K}]}{[8.7 \text{ dm}^3][273 \text{ K}]} = 179.6 \text{ KPa}$$

4. A helium balloon is fully inflated at 1.2 L. When the clerk is filling the balloon, she stops to make sure it is not going to explode and checks the pressure. The pressure is 5.7 KPa and the temperature at the store is 24°C. The balloon is 0.75L full. She stops when the balloon is fully inflated and the pressure is 7.2 KPa. When the customer takes the balloon outside, it explodes. What was the temperature outdoors?

$$T = \frac{[24^\circ\text{C} + 273\text{K}][1.2 \text{ L}][7.2 \text{ KPa}]}{[0.75 \text{ L}][5.7 \text{ KPa}]} = 600 \text{ K}$$

5. Oxygen gas is added to a rebreather at the same rate as it would be in the open air. The volume is 4.8 dm<sup>3</sup> and the temperature is 25 °C. The pressure is 1 atm. Once the tank is lowered under the water the pressure increases to 2 atm and the temperature drops to 12°C. Is the tank in danger of exploding if it's maximum volume is 5.0 dm<sup>3</sup>?

$$V_2 = \frac{[4.8 \text{ dm}^3][1 \text{ atm}][12^\circ\text{C} + 273\text{K}]}{[2 \text{ atm}][25^\circ\text{C} + 273\text{K}]} = 2.1 \text{ dm}^3, \text{ not in danger of exploding}$$

6. Ozone is formed due to electricity passing through the air and splitting some O<sub>2</sub> atoms which then join briefly to form O<sub>3</sub>, ozone. Your olfactory sense can detect 1.00 mL ozone in 10 L of air at STP. You are surprised that the noreaster is accompanied by

# Gas Laws And Scuba Diving Worksheet Answer

**William Maurice Dehn**



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