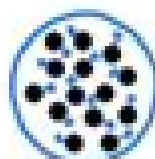




- This is colder because it is more condensed and so there is less movement.



- | | | | |
|---------------------|-------------------------------------|-------------------------------------|-----------------|
| Robert Boyle | Jacques Charles | Guillaume Amontons | Amedeo Avogadro |
| $P_1 V_1 = P_2 V_2$ | $\frac{V_1}{T_1} = \frac{V_2}{T_2}$ | $\frac{P_1}{T_1} = \frac{P_2}{T_2}$ | $V \propto n$ |

Chemists combine all of the relationships seen in Models 1 and 2 into one law—the **Ideal Gas Law**. It is one equation that describes gas behavior and the relationship among all four variables, P , V , T , and n . In the Ideal Gas Law the proportionality constant is represented by the letter R (rather than the generic k).

- PTV = Rm

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