

**The Gas Laws – Ch. 10** (p.313-322)

1. The gas left in a used aerosol can is at a pressure of 1 atm at 27°C. If this can is thrown into a fire, what is the internal pressure of the gas when its temperature reaches 927°C?

GIVEN	GAS LAW	WORK
T1= 300K P1= 101 kPa T2= 1200K P2= x	BOYLE	101 / 300 = x / 1200 x = 404
	FORMULA	
	P1/T1=P2/T2	
ANSWER: 404 kPa OR 4 ATM		

2. A sample of carbon dioxide occupies a volume of 3.50 L at 125 kPa. What pressure would the gas exert if the volume were decreased to 2.00 L?

GIVEN	GAS LAW	WORK
V1 = 3.50 L P1 = 125 kPa V2 = 2.00 L P2 = x	CHARLES	(125) (3.50) = (2.00) (x) x = 218.75 kPa
	FORMULA	
	P1V1 = P2V2	
ANSWER: 218.75 kPa		

3. A sample of propane occupies 250.0 L at 125 kPa and 38°C. Find its volume at 100.0 kPa and 95°C.

GIVEN	GAS LAW	WORK
v1 = 250.0 L P1 = 125 kPa T1 = 311 K V2 = x P2 = 100.0 kPa T2 = 368 K	COMBINED	(125) (250.0) / 311 = (100.0) (x) / 368 x = 369.77
	FORMULA	
	P1v1/T1=P2v2/T2	
ANSWER: 369.77 L		

# Gas Law Worksheet 2 Answers

**Michael Brown**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## Gas Law Worksheet 2 Answers:

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## **Table of Contents Gas Law Worksheet 2 Answers**

1. Understanding the eBook Gas Law Worksheet 2 Answers
  - The Rise of Digital Reading Gas Law Worksheet 2 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Law Worksheet 2 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Law Worksheet 2 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Law Worksheet 2 Answers



- Personalized Recommendations
  - Gas Law Worksheet 2 Answers User Reviews and Ratings
  - Gas Law Worksheet 2 Answers and Bestseller Lists
5. Accessing Gas Law Worksheet 2 Answers Free and Paid eBooks
    - Gas Law Worksheet 2 Answers Public Domain eBooks
    - Gas Law Worksheet 2 Answers eBook Subscription Services
    - Gas Law Worksheet 2 Answers Budget-Friendly Options
  6. Navigating Gas Law Worksheet 2 Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Gas Law Worksheet 2 Answers Compatibility with Devices
    - Gas Law Worksheet 2 Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Gas Law Worksheet 2 Answers
    - Highlighting and Note-Taking Gas Law Worksheet 2 Answers
    - Interactive Elements Gas Law Worksheet 2 Answers
  8. Staying Engaged with Gas Law Worksheet 2 Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Gas Law Worksheet 2 Answers
  9. Balancing eBooks and Physical Books Gas Law Worksheet 2 Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Gas Law Worksheet 2 Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Gas Law Worksheet 2 Answers
    - Setting Reading Goals Gas Law Worksheet 2 Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Gas Law Worksheet 2 Answers

- Fact-Checking eBook Content of Gas Law Worksheet 2 Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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