

Empirical and Molecular Formulas Practice

Read the question carefully for the data provided. Answers must be shown for credit.

Find empirical formula:

1. A compound is found to contain 40.0 g of carbon and 6.67 g of hydrogen. Find the empirical formula.

$$40.0 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 3.33 \text{ mol C}$$

$$6.67 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 6.60 \text{ mol H}$$

C₃H₆

2. A compound is found to contain 20.0 g of carbon, 26.7 g of oxygen, and 6.67 g of hydrogen. Find the empirical formula.

$$20.0 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.67 \text{ mol C}$$

$$26.7 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 1.67 \text{ mol O}$$



$$6.67 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 6.60 \text{ mol H} \quad \frac{6.60 \text{ mol H}}{1.67 \text{ mol C}} = 4 \text{ H}$$

3. A compound is found to contain 40.0 g of carbon, 6.67 g of hydrogen, and 53.3 g of oxygen. Find the empirical formula.

$$40.0 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 3.33 \text{ mol C}$$



$$6.67 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 6.60 \text{ mol H} \quad \frac{6.60 \text{ mol H}}{3.33 \text{ mol C}} = 2 \text{ H}$$

$$53.3 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 3.33 \text{ mol O} \quad \frac{3.33 \text{ mol O}}{3.33 \text{ mol C}} = 1 \text{ O}$$

4. A 25.00 g compound is analyzed and found to contain 10.00 g carbon, 1.34 g hydrogen, and 13.66 g oxygen. Find the empirical formula.

$$10.00 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 0.833 \text{ mol C}$$



$$1.34 \text{ g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 1.33 \text{ mol H}$$

$$13.66 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.854 \text{ mol O} \quad \frac{1.33 \text{ mol H}}{0.854 \text{ mol O}} = 1.56 \text{ H}$$

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