

Example 1: Calculate the empirical formula for a substance containing 1.11 g of Se and 1.61 g of F.

$$1.11 \text{ g } \cancel{\text{Se}} \times \frac{1 \text{ mol Se}}{78.971 \text{ g } \cancel{\text{Se}}} = 0.0141 \text{ mol Se}$$

$$1.61 \text{ g } \cancel{\text{F}} \times \frac{1 \text{ mol F}}{18.998 \text{ g } \cancel{\text{F}}} = 0.0847 \text{ mol F}$$

$$\text{Se} \frac{0.0141}{0.0141} \text{F} \frac{0.0847}{0.0141} = \text{SeF}_6$$

Empirical Formula Answers Only

VM Jensen

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the VM Jensen text.

Empirical Formula Answers Only:

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