

Empirical and Molecular Formula Worksheet ANSWER KEY

Write the empirical formula for the following compounds.

- 1) C_6H_6 **CH**
- 6) C_8H_{14} **C_4H_7**
- 7) WO_2 **WO_2**
- 8) $C_2H_6O_2$ **CH_3O**
- 9) $X_{10}Y_{12}$ **X_5Y_6**
- 6) A compound with an empirical formula of C_2OH_4 and a molar mass of 88 grams per mole. What is the molecular formula of this compound? **$C_8O_4H_{16}$**
- 7) A compound with an empirical formula of C_4H_4O and a molar mass of 136 grams per mole. What is the molecular formula of this compound? **$C_8H_8O_2$**
- 8) A compound with an empirical formula of $CFBrO$ and a molar mass of 253.8 grams per mole. What is the molecular formula of this compound? **$C_2F_2Br_2O_2$**
- 9) A compound with an empirical formula of C_3H_3N and a molar mass of 46 grams per mole. What is the molecular formula of this compound? **$C_6H_6N_2$**
- 10) A well-known reagent in analytical chemistry, dimethylglyoxime, has the empirical formula C_2H_5NO . If its molar mass is 116.1 g/mol, what is the molecular formula of the compound? **$C_8H_{10}N_2O_2$**
12. A certain blue solid contains 36.84% N and 63.16% O. What is the empirical formula of this compound? **The ratios are $N_{1.50}O_{1.50}$. Since 1.50 is not close to a whole number, we multiply *both* subscripts by 2. The empirical formula is thus N_2O_2 . (The name is dinitrogen trioxide.)**
13. A sample of indium chloride weighing 0.5000 g is found to contain 0.2404 g of chlorine. What is the empirical formula of the indium compound? **$InCl_2$**
14. An unknown compound was found to have a percent composition as follows: 47.0 % potassium, 14.5 % carbon, and 38.5 % oxygen. What is its empirical formula? If the true molar mass of the compound is 166.22 g/mol, what is its molecular formula? **$K_2C_3O_4$**
15. Rubbing alcohol was found to contain 60.0 % carbon, 13.4 % hydrogen, and the remaining mass was due to oxygen. What is the empirical formula of rubbing alcohol? **C_3H_8O**

Empirical And Molecular Formula Answer

Eileen Ramsden



Empirical And Molecular Formula Answer:

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



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