

## Chemistry Honors Worksheet: Formulas of a Compound

1.
  - a.) moles The subscripts in a chemical formula represent the ratio of ( moles / mass ) in a compound
  - b.) CO Which compound has the greatest percentage of carbon? (CO, CO<sub>2</sub>, PbCO<sub>3</sub> )
  - c.) CH<sub>4</sub> What is the empirical formula of a compound that contains 0.2 moles of C and 0.8 moles of H
  - d.) 83 g/mol Calculate the molecular mass of a compound if 0.02461 moles has a mass of 2.05 g.
  - e.) Same The % composition of an empirical formula is ( more / less / the same ) than/as its molecular formula.
  - f.) one How many moles of V are indicated in one mole of the compound V(C<sub>2</sub>O<sub>4</sub>)<sub>2</sub>?
  - g.)  $1 \times 10^{24}$  atoms A beaker of water contains  $2 \times 10^{24}$  atoms of H, how many atoms of O are also in the beaker?
  - h.)  $6.02 \times 10^{23}$  1 mole of a molecular compound has a mass of 165 grams. How many molecules are there?
  - i.) 0% Fool's Gold is iron pyrite, FeS. What is the percentage of gold in a 1.0 gram sample of fool's gold?

2. The pesticide DDT has been banned in the United States since 1962. This compound has a formula C<sub>14</sub>H<sub>9</sub>Cl<sub>5</sub>. Answer all questions referring to DDT. Note: Use 35.5 g/mol for Cl 709.5 g/mol

- a.) molecular Is this compound ionic or molecular?
- b.)  $5.2 \times 10^{15}$  How many molecules of DDT are in a  $6.1 \times 10^{-4}$  g sample?
- c.)  $1.53 \times 10^{17}$  atoms H How many atoms of H are in 2.00 mg of DDT?
- d.) 7.51 g Cl How many grams of Cl are contained in 10.0 grams of DDT?
- e.)  $7.84 \times 10^{21}$  atoms C If a sample  $5.04 \times 10^{21}$  atoms of H, how many atoms of C would there also be?
- f.)  $2.97 \times 10^{20}$  atoms C How many atoms of C are in a 25.0 mg of DDT?

3. A compound of tantalum, Ta, and oxygen was produced in a lab by heating tantalum in a crucible. The data was collected.

Mass of crucible:	24.22 g	> 1.72 g Ta	0.0106 mol Ta
Mass of crucible and tantalum:	26.14 g		
Mass of crucible and tantalum oxide:	26.96 g	> 0.42 g O	0.02625 mol O

- a.) Calculate the percent composition of the compound. 82.1% Ta 17.9% O
- b.) What is the empirical formula of this compound? TaO<sub>2.5</sub> - Ta<sub>2</sub>O<sub>5</sub>
- c.) Is this compound ionic or covalent? ionic
- d.) Name this compound: Tantalum(V) oxide \* Need Roman Numerals!!!

4. Determine the empirical formula of the following compounds:

- a.) 63.0g Rb, 5.90g O Rb<sub>2</sub>O
- b.) 4.95mg Th, 1.37mg S ThS<sub>2</sub>
- c.) 2.13g Na, 2.32g As, 1.98g O Na<sub>3</sub>AsO<sub>4</sub>
- d.) 32.8% Cr, 67.2% Cl CrCl<sub>3</sub>
- e.) 58.0% Rb, 9.50% N, 32.5% O RbNO<sub>3</sub>

# Empirical Molecular Formula Answers

**Arun Bahl & B S Bahl**



## **Empirical Molecular Formula Answers:**

Understanding General Chemistry Atef Korchef, 2022-03-07 Understanding General Chemistry details the fundamentals of general chemistry through a wide range of topics relating the structure of atoms and molecules to the properties of matter. Written in an easy to understand format with helpful pedagogy to fuel learning, the book features main objectives at the beginning of each chapter, get smart sections, and check your reading section at the end of each chapter. The text is filled with examples and practices that illustrate the concepts at hand. In addition, a summary and extensive MCQs, exercises, and problems with the corresponding answers and explanations are readily available. Additional features include Alerts students to common mistakes and explains in simple ways and clear applications how to avoid these mistakes. Offers answers and comments alongside sample problems enabling students to self-evaluate their skill level. Includes powerful methods, easy steps, simple and accurate interpretations, and engaging applications to help students understand complex principles. Provides a bridge to more complex topics such as solid state chemistry, organometallic chemistry, chemistry of main group elements, inorganic chemistry, and physical chemistry. This introductory textbook is ideal for chemistry courses for non-science majors as well as health sciences and preparatory engineering students. **E3 Chemistry Review Book - 2018 Home Edition**

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& B S Bahl,2019 With an increased focus on fundamentals this new edition of A Textbook of Organic Chemistry continues to present the time tested functional group approach to the subject This examination oriented book breaks the intricacies of Organic Chemistry into easy to understand steps which gives the student the necessary foundation to build upon learn and understand Organic Chemistry in a way that is efficient as well as long lasting      Calculations for GCSE Chemistry Eileen Ramsden,2002 This fully revised edition is in line with the revised 2002 National Curriculum requirements and focuses on quantitative chemistry in science Written to match all major GCSE specifications the text covers all types of numerical questions from first principles For each topic a concise treatment of the underlying theory is followed by problems grouped into three sections of increasing difficulty Calculations based on round number molar masses are included to enable students to concentrate on the chemical basis of the problems rather than arithmetical manipulation      **Chemistry** Eileen Ramsden,2001 This third edition of Key Science Chemistry has been fully revised to meet the requirements of all 2001 GCSE specifications It is aimed at middle ability students but contains enough material for high achievers Topics are clearly differentiated between core material for GCSE science Double Award Single Award and extension material for GCSE science chemistry      Basics for Chemistry David A. Ucko,2013-09-24 Basics of Chemistry provides the tools needed in the study of General Chemistry such as problem solving skills calculation methods and the language and basic concepts of chemistry The book is designed to meet the specific needs of underprepared students Concepts are presented only as they are needed and developed from the simple to the complex The text is divided into 18 chapters each covering some particular aspect of chemistry such as matter energy and measurement the properties of atoms description of chemical bonding study of chemical change and nuclear and organic chemistry Undergraduate students will find the book as a very valuable academic material      Chemistry Richard Post,Chad Snyder,Clifford C. Houk,2020-08-20 A practical complete and easy to use guide for understanding major chemistry concepts and terms Master the fundamentals of chemistry with this fast and easy guide Chemistry is a fundamental science that touches all other sciences including biology physics electronics environmental studies astronomy and more Thousands of students have successfully used the previous editions of Chemistry Concepts and Problems A Self Teaching Guide to learn chemistry either independently as a refresher or in parallel with a college chemistry course This newly revised edition includes updates and additions to improve your success in learning chemistry This book uses an interactive self teaching method including frequent questions and study problems increasing both the speed of learning and retention Monitor your progress with self tests and master chemistry quickly This revised Third Edition provides a fresh step by step approach to learning that requires no prerequisites lets you work at your own pace and reinforces what you learn ensuring lifelong mastery Master the science of basic chemistry with this innovative self paced study guide Teach yourself chemistry refresh your knowledge in preparation for medical studies or other coursework or enhance your college chemistry course Use self study features including review questions and quizzes to ensure that you re really learning the

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**An Introduction to Chemistry** Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

*Counting Moles* Nigel P. Freestone, 2016-01-04 Students studying chemistry often struggle with the mole Counting Moles provides an effective aid to learning by giving clear and confident presentation of the essentials of the mole concept needed by those starting chemistry courses This user friendly self teach e book is split into six chapters which sequentially introduce the mole calculating frame to help solve problems Over 200 fully worked examples are given along with several hundred questions The mole concept is applied to topics such as relative atomic mass and relative formula mass percentage composition empirical and molecular formula The book also covers concentration its units volumetric analysis and the relationship between volume mass and moles of gases Counting Moles culminates in you taking a Mole Driving Test On passing this test you are issued with a Counting Moles Driving License that will give you all the confidence required to correctly answer all mole calculations

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**Fundamentals of Chemistry: A Modern Introduction (1966)** Frank Brescia,2012-12-02 Fundamentals of Chemistry A Modern Introduction focuses on the formulas processes and methodologies used in the study of chemistry The book first looks at general and historical remarks definitions of chemical terms and the classification of matter and states of aggregation The text then discusses gases Ideal gases pressure of a gas confined by a liquid Avogadro's Law and Graham's Law are described The book also discusses aggregated states of matter atoms and molecules chemical equations and arithmetic thermochemistry and chemical periodicity The text also highlights the electronic structures of atoms Quantization of electricity spectra of elements quantization of the energy of an electron associated with nucleus the Rutherford Bohr nuclear theory hydrogen atom and representation of the shapes of atomic orbitals are explained The text also highlights the types of chemical bonds hydrocarbons and their derivatives intermolecular forces solutions and chemical equilibrium The book focuses as well on ionic solutions galvanic cells and acids and bases It also discusses the structure and basicity of hydrides and oxides The reactivity of hydrides charge of dispersal and basicity effect of anionic charge inductive effect and basicity and preparation of acids are described The book is a good source of information for readers wanting to study chemistry

**Educart ICSE Class 10 Question Bank 2025 Chemistry One Shot for 2024-25 Exam** Educart,Sir Tarun Rupani,2024-06-17 What You Get Analytical based Q's All Important Q's Educart ICSE Class 10 CHEMISTRY One Shot Question Bank 2024-25 Updated for 2025 Exam Strictly Based on ICSE Specimen Paper 15th May 2024 Competency based Q's as per revised ICSE pattern Push Yourself to perform well with High Order Q's Direct All Types of Memory based Q's and PYQ's Why choose this book First Book to introduce 25% High Order Ability Questions as per ICSE Specimen Papers 2025

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