

Question 13
Acids and Bases IV
Answered

13. A buffer solution consisting of H_2PO_4^- is added to a 1.0 L solution containing H_2PO_4^- . How the concentration of H_2PO_4^- changes. Assume that the total PO_4^{3-} family stays constant. Assume that the overall volume of the solution does not change.

HPO_4^{2-} and H_2PO_4^- is H_2PO_4^-
 And H_2PO_4^- and H_2PO_4^- is H_2PO_4^-
 And H_2PO_4^- will decrease completely. The addition of H_2PO_4^- ions will shift the equilibrium of the reaction to the left, according to Le Chatelier's principle. Because H_2PO_4^- will combine with H_2PO_4^- to form the products H_2PO_4^- . This will continue until the product of $[\text{H}_2\text{PO}_4^-][\text{H}_2\text{PO}_4^-]$ equals K_a once again. This will decrease the concentration of H_2PO_4^- in the solution.

14. A buffer solution is created by dissolving 10.0 moles of weak acid, and 10.0 moles of its conjugate base in 1.0 L of water. Assume that the volume of the solution does not change.
 a. What the pH of the solution

15. Consider a solution with $\text{pH} = 8.00$. $\text{pH} = \text{pH}_a + \log\left(\frac{[\text{A}^-]}{[\text{HA}]}\right) = -\log(1.0 \times 10^{-7}) + \log\left(\frac{[\text{A}^-]}{[\text{HA}]}\right) = 8.00$
 so $\log\left(\frac{[\text{A}^-]}{[\text{HA}]}\right) = 8.00 + \log(1.0 \times 10^{-7}) = 8.00 - 7.00 = 1.00$
 $\frac{[\text{A}^-]}{[\text{HA}]} = 10^1 = 10$

b. What the pH after 0.25 moles of HCl is added to the solution. Assume that the volume does not change.

16. The solution is initially 0.10 M H_2PO_4^- and 0.10 M H_2PO_4^-

	H_2PO_4^-	\rightleftharpoons	H_2PO_4^-	\rightleftharpoons	H_2PO_4^-
Initial	0.10 mol/L		0.10 mol/L		0.10 mol/L
Change	-0.025 mol/L		-0.025 mol/L		+0.025 mol/L
Equilibrium	0.075 mol/L		0.075 mol/L		0.125 mol/L

$\text{pH} = \text{pH}_a + \log\left(\frac{[\text{A}^-]}{[\text{HA}]}\right) = -\log(6.2 \times 10^{-8}) + \log\left(\frac{0.125}{0.075}\right) = 7.20$

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