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Chemistry 12  
 Worksheet 4-2

82

## Bronsted Acids and Equilibria

1. Write the formula for a *proton* (1 mark)  $H^+$
2. Write the formula for a *hydrated proton* (1 mark)  $H_3O^+$
3. Write the formula for a *hydronium* ion (1 mark)  $H_3O^+$
4. Give the *Arrhenius* definition of an *acid* (1 mark) any subst. which releases  $H^+$  ions in water
5. Give the *Arrhenius* definition of a *base* (1 mark) any subst. which releases  $OH^-$  ions in water.
6. Give the *Bronsted* definition of an *acid* (1 mark) a species which donates protons to another species
7. Give the *Bronsted* definition of a *base* (1 mark) a species which accepts protons from another species
8. Given the equation:  $HCO_3^- + H_2S \rightleftharpoons H_2CO_3 + HS^-$ 
  - a) The **acid** on the left side is (1 mark)  $H_2S$
  - b) The **base** on the left side is (1 mark)  $HCO_3^-$
  - c) The **acid** on the right side is (1 mark)  $H_2CO_3$
  - d) The **base** on the right side is (1 mark)  $HS^-$

# Lecture 30 Acids And Bases Solutions Worksheet

**Robert A. Richardson**



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