

**Exercice 16**  
**Autonomous Work (20 minutes)**

13. A 1.00 mol sample of  $\text{NH}_3(\text{g})$  is added to water. The final volume of the solution is 1.0 L.
- a. Calculate the pH of the solution.

$$\begin{aligned} \text{NH}_3(\text{g}) + \text{H}_2\text{O}(\text{l}) &\rightleftharpoons \text{NH}_4^+(\text{aq}) + \text{OH}^-(\text{aq}) \\ K_b &= 1.8 \times 10^{-5} \\ \text{pH} &= 11.25 \end{aligned}$$

- b. Is the solution acidic or basic? Explain.

The solution is basic, as the pH is less than 7.

14. Calculate the pH of a 0.10 M solution of  $\text{HNO}_2$  ( $K_a = 4.5 \times 10^{-4}$  at 25°C).

|  |  |  |  |                                                                                 |
|--|--|--|--|---------------------------------------------------------------------------------|
|  |  |  |  | $K_a = \frac{[\text{H}^+][\text{NO}_2^-]}{[\text{HNO}_2]} = 4.5 \times 10^{-4}$ |
|  |  |  |  | $[\text{H}^+] = [\text{NO}_2^-] = x$                                            |
|  |  |  |  | $[\text{HNO}_2] = 0.10 - x$                                                     |
|  |  |  |  | $4.5 \times 10^{-4} = \frac{x^2}{0.10 - x}$                                     |
|  |  |  |  | $x = 6.7 \times 10^{-3} \text{ M}$                                              |
|  |  |  |  | $\text{pH} = -\log(6.7 \times 10^{-3}) = 2.17$                                  |
|  |  |  |  | $\text{pH} = 2.17$                                                              |

15. Calculate the equilibrium concentration of a solution made up of 0.10 M  $\text{HNO}_2$ .

|                                               |                                                                                                                                               |
|-----------------------------------------------|-----------------------------------------------------------------------------------------------------------------------------------------------|
| $\text{pH} = 2.17$                            | $K_a = 4.5 \times 10^{-4}$                                                                                                                    |
| $[\text{H}^+] = 6.7 \times 10^{-3} \text{ M}$ | $[\text{HNO}_2] = \frac{[\text{H}^+][\text{NO}_2^-]}{K_a} = \frac{(6.7 \times 10^{-3})^2}{4.5 \times 10^{-4}} = 9.9 \times 10^{-2} \text{ M}$ |
| $\text{pH} = 2.17$                            | $[\text{HNO}_2] = 9.9 \times 10^{-2} \text{ M}$                                                                                               |

# Lecture 30 Acids And Bases I

**M Tight**



### Lecture 30 Acids And Bases I:

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